

Answers

Chapter 1

Exercises

- 1 (a) $\text{CuCO}_3 \rightarrow \text{CuO} + \text{CO}_2$
 (b) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
 (c) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 (d) $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 (e) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- 2 (a) $2\text{K} + 2\text{H}_2\text{O} \rightarrow 2\text{KOH} + \text{H}_2$
 (b) $\text{C}_2\text{H}_5\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$
 (c) $\text{Cl}_2 + 2\text{KI} \rightarrow 2\text{KCl} + \text{I}_2$
 (d) $4\text{CrO}_2 \rightarrow 2\text{Cr}_2\text{O}_3 + 3\text{O}_2$
 (e) $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 3\text{CO} + 2\text{Fe}$
- 3 (a) $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 (b) $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$
 (c) $3\text{Cu} + 8\text{HNO}_3 \rightarrow 3\text{Cu}(\text{NO}_3)_2 + 2\text{NO} + 4\text{H}_2\text{O}$
 (d) $6\text{H}_2\text{O}_2 + 2\text{N}_2\text{H}_4 \rightarrow 2\text{N}_2 + 10\text{H}_2\text{O} + \text{O}_2$
 (e) $4\text{C}_3\text{H}_7\text{N} + 15\text{O}_2 \rightarrow 8\text{CO}_2 + 14\text{H}_2\text{O} + 2\text{N}_2$
- 4 (a) Sand and water: heterogeneous
 (b) Smoke: heterogeneous
 (c) Sugar and water: homogeneous
 (d) Salt and iron filings: heterogeneous
 (e) Ethanol and water: homogeneous
 (f) Steel: homogeneous
- 5 (a) $2\text{KNO}_3(\text{s}) \rightarrow 2\text{KNO}_2(\text{s}) + \text{O}_2(\text{g})$
 (b) $\text{CaCO}_3(\text{s}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{CaSO}_4(\text{s}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$
 (c) $2\text{Li}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{LiOH}(\text{aq}) + \text{H}_2(\text{g})$
 (d) $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{NaCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$
 (e) $2\text{C}_2\text{H}_6(\text{g}) + 9\text{O}_2(\text{g}) \rightarrow 6\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{l})$

In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

- 6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)
- 7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.
- 8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.
- 9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.
- 10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E–D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.
- 11 B
- 12 AWW 01_EX_12

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Clifford Lane



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1. Understanding the eBook Pearson Chemistry Study Guide Answers
 - The Rise of Digital Reading Pearson Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Study Guide Answers
 - Personalized Recommendations
 - Pearson Chemistry Study Guide Answers User Reviews and Ratings
 - Pearson Chemistry Study Guide Answers and Bestseller Lists

5. Accessing Pearson Chemistry Study Guide Answers Free and Paid eBooks
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 - Pearson Chemistry Study Guide Answers Compatibility with Devices
 - Pearson Chemistry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Study Guide Answers
 - Highlighting and Note-Taking Pearson Chemistry Study Guide Answers
 - Interactive Elements Pearson Chemistry Study Guide Answers
8. Staying Engaged with Pearson Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Pearson Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Study Guide Answers
 - Setting Reading Goals Pearson Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Pearson Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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