#### ENGAGE

Why does the atomic number but not the mass number change in positron emission?

#### SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission

#### TRY IT FIRST

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

#### SOLUTION GUIDE

ANALYZE THE PROBLEM	Given	Need	Connect
	Mn-49, positron	balanced nuclear	mass number, atomic
	emission	equation	number of new nucleus

STEP 1 Write the incomplete nuclear equation.

$$^{49}_{29}Mm \longrightarrow 7 + ^{9}_{47}e$$

STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = 7 + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49$$
 (mass number of new nucleus)

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = 7 + 1$$

$$25 - 1 = 7$$

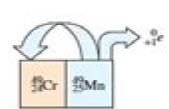
STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written 50r.

SIEP 5 Complete the nuclear equation.

#### STUDY CHECK S.4.

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

#### ANSWER



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Puja Mehta

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- 1. Understanding the eBook Pearson Chemistry Practice Problem Answers
  - The Rise of Digital Reading Pearson Chemistry Practice Problem Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pearson Chemistry Practice Problem Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Practice Problem Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pearson Chemistry Practice Problem Answers

- Personalized Recommendations
- Pearson Chemistry Practice Problem Answers User Reviews and Ratings
- Pearson Chemistry Practice Problem Answers and Bestseller Lists
- 5. Accessing Pearson Chemistry Practice Problem Answers Free and Paid eBooks
  - Pearson Chemistry Practice Problem Answers Public Domain eBooks
  - Pearson Chemistry Practice Problem Answers eBook Subscription Services
  - Pearson Chemistry Practice Problem Answers Budget-Friendly Options
- 6. Navigating Pearson Chemistry Practice Problem Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Pearson Chemistry Practice Problem Answers Compatibility with Devices
  - Pearson Chemistry Practice Problem Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Chemistry Practice Problem Answers
  - Highlighting and Note-Taking Pearson Chemistry Practice Problem Answers
  - Interactive Elements Pearson Chemistry Practice Problem Answers
- 8. Staying Engaged with Pearson Chemistry Practice Problem Answers
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chemistry Practice Problem Answers
- 9. Balancing eBooks and Physical Books Pearson Chemistry Practice Problem Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chemistry Practice Problem Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Practice Problem Answers
  - Setting Reading Goals Pearson Chemistry Practice Problem Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Practice Problem Answers

- Fact-Checking eBook Content of Pearson Chemistry Practice Problem Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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