



Learning Objectives

- Become skilled in the technique of acid-base titration, including the standardization of a base solution.
- Be able to end a titration at the precise endpoint.
- Determine the molecular weight of an unknown acid by titrating with a standardized base solution.

- 100g H_3PO_4
- 99.0g mass = 99.0g/mol
- 3% 0.01L NaOH
- M of NaOH = ?

When you
titrating the
12.80g of the
11.80g



(x3)

11

-3



acid + base → water + salt



$$100g H_3PO_4 \left(\frac{1 \text{ mol } H_3PO_4}{99.0g H_3PO_4} \right) \left(\frac{3 \text{ mol } NaOH}{1 \text{ mol } H_3PO_4} \right) = 0.003 \text{ mol } NaOH$$

12.80g of the
11.80g

$$M = \frac{mm}{L}$$

$$M = \frac{0.003}{0.035} = 0.8658 = 0.866$$

$$\begin{array}{r} 25.60 \text{ mL} \\ - 2.21 \text{ mL} \\ \hline 23.39 \text{ mL} \end{array}$$

$$\begin{array}{r} 25 \\ - 2 \\ \hline 23 \end{array}$$



Pearson Chemistry Lab Answers

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