

AP CHEMISTRY CHAPTER REVIEW

CHAPTER 7: COVALENT BONDING

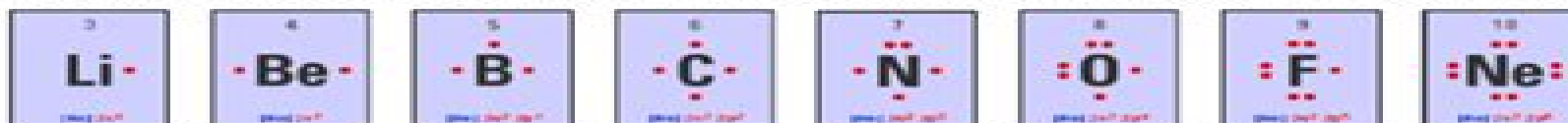
- You should understand the nature of the covalent bond.
- You should be able to draw the Lewis electron-dot structure for any atom, molecule, or polyatomic ion.
- When a single Lewis structure does not adequately reflect the properties of a substance, you should be able to draw resonance forms for it.
- You should be able to calculate the formal charge on an atom in an electron-dot structure.
- You should be able to give examples of substances that violate the octet rule, either because they are electron-deficient or because they have expanded octets.
- You should be able to apply the principles of VSEPR theory to identify the geometry of any molecule or ion.
- You should be able to predict the polarity of a substance based on its geometry.
- You should be able to identify the orbital hybridization of the central atom in a molecule or polyatomic ion.
- You should be able to identify the sigma and pi bonds in an electron-dot structure.

7.1 LEWIS STRUCTURES: THE OCTET RULE

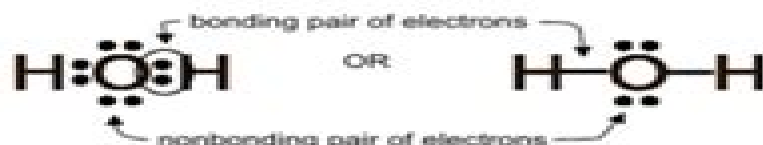
Noble gases are generally nonreactive because they have a full valence shell. Many atoms form monatomic ions by either **losing electrons** or **gaining electrons**, in order to achieve the electron configuration of a noble gas. Gilbert Lewis suggested that nonmetal atoms can acquire a noble gas configuration by **sharing electrons** to form a **covalent bond**.



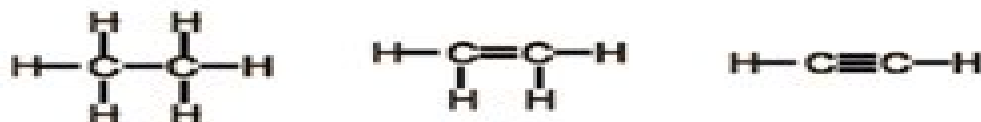
The Lewis electron dot structures below indicate the valence electrons for elements in Groups 1-2 and Groups 13-18.



In a Lewis electron dot structure, the electrons normally exist in pairs. **Bonding pairs** are covalent bonds. **Nonbonding pairs** are lone pairs, or unshared pairs.



A single bond consists of one bonding pair. A double bond consists of two bonding pairs. A triple bond consists of three bonding pairs.



All other things being equal, triple bonds are shorter than double bonds, which are shorter than single bonds.



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R Barnett



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