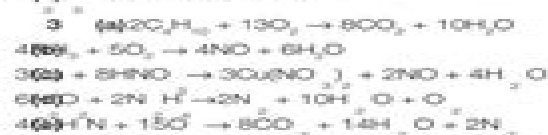
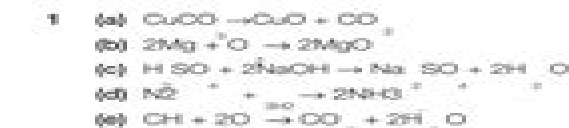


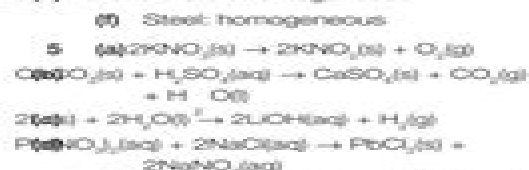
Answers

Chapter 1

Exercises



4 (a) Sand and water: heterogeneous
 (b) Sugar and water: homogeneous
 (c) Oil and iron filings: heterogeneous
 (d) Endol and water: homogeneous



In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)

7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.

8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.

9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.

10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E-D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.

11 B

12 AW 01_EX_12

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1. Understanding the eBook Pearson Chemistry Chapters 5 Assessment Answers
 - The Rise of Digital Reading Pearson Chemistry Chapters 5 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Chapters 5 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Chapters 5 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Chapters 5 Assessment Answers
 - Personalized Recommendations
 - Pearson Chemistry Chapters 5 Assessment Answers User Reviews and Ratings
 - Pearson Chemistry Chapters 5 Assessment Answers and Bestseller Lists

5. Accessing Pearson Chemistry Chapters 5 Assessment Answers Free and Paid eBooks
 - Pearson Chemistry Chapters 5 Assessment Answers Public Domain eBooks
 - Pearson Chemistry Chapters 5 Assessment Answers eBook Subscription Services
 - Pearson Chemistry Chapters 5 Assessment Answers Budget-Friendly Options
6. Navigating Pearson Chemistry Chapters 5 Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Chapters 5 Assessment Answers Compatibility with Devices
 - Pearson Chemistry Chapters 5 Assessment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Chapters 5 Assessment Answers
 - Highlighting and Note-Taking Pearson Chemistry Chapters 5 Assessment Answers
 - Interactive Elements Pearson Chemistry Chapters 5 Assessment Answers
8. Staying Engaged with Pearson Chemistry Chapters 5 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Chapters 5 Assessment Answers
9. Balancing eBooks and Physical Books Pearson Chemistry Chapters 5 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Chapters 5 Assessment Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Chapters 5 Assessment Answers
 - Setting Reading Goals Pearson Chemistry Chapters 5 Assessment Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Chapters 5 Assessment Answers
 - Fact-Checking eBook Content of Pearson Chemistry Chapters 5 Assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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