

REVIEW QUESTIONS

- Identify the reactant that acts as an acid in each of the following reactions.
 - $\text{NH}_4^+(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{NH}_3(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$
 - $\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$
 - $\text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_3^{2-}(\text{aq})$
 - $\text{SO}_4^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq}) \rightarrow \text{HSO}_4^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$
 - $\text{CO}_3^{2-}(\text{aq}) + \text{CH}_3\text{COOH}(\text{aq}) \rightarrow \text{HCO}_3^-(\text{aq}) + \text{CH}_3\text{COO}^-(\text{aq})$
- Write balanced equations to show that in water:
 - PO_4^{3-} acts as a base
 - H_2PO_4^- acts as an amphiprotic substance
 - H_2S acts as an acid.
- Write the formula for the conjugate of:
 - the acid HCl
 - the base OH^-
 - the base O^{2-}
 - HSO_4^- when it acts as an acid.
- Using suitable examples, distinguish between:
 - a diprotic and an amphiprotic substance
 - a strong and a weak acid.
- Draw a structural formula of the monoprotic ethanoic acid molecule. Identify which proton is donated in an
- Which of the following species is both diprotic and amphiprotic?
 - HCl
 - HSO_4^-
 - H_2SO_3
 - H_2PO_4^-
- Complete and balance the following reactions:
 - $\text{Fe}(\text{s}) + 2\text{HF}(\text{aq}) \rightarrow$
 - $\text{HClO}_4(\text{aq}) + \text{LiOH}(\text{aq}) \rightarrow$
 - $\text{HNO}_3(\text{aq}) + \text{KHCO}_3(\text{aq}) \rightarrow$
 - $\text{Li}_2\text{CO}_3(\text{aq}) + 2\text{CH}_3\text{COOH}(\text{aq}) \rightarrow$
- For each of the chemical equations in Question 10, write the corresponding net ionic equation.
- What volume of $0.0010 \text{ mol L}^{-1}$ potassium hydroxide

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