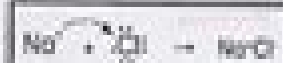


IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. $\text{K} + \text{F}$



2. $\text{Mg} + \text{I}$



3. $\text{Be} + \text{S}$



4. $\text{Na} + \text{O}$



5. $\text{Al} + \text{Br}$



COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:



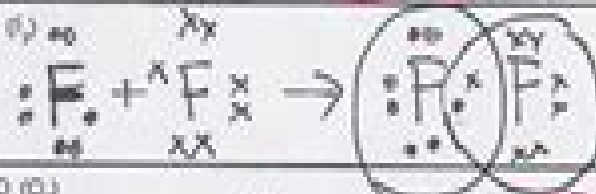
Note that hydrogen is content with 2, not 8, electrons.

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. $\text{H} + \text{H} (\text{H}_2)$



2. $\text{F} + \text{F} (\text{F}_2)$



3. $\text{O} + \text{O} (\text{O}_2)$



4. $\text{N} + \text{N} (\text{N}_2)$



5. $\text{C} + \text{O} (\text{CO})$



6. $\text{H} + \text{O} (\text{H}_2\text{O})$



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LP Steffe



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1. Understanding the eBook Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - The Rise of Digital Reading Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Personalized Recommendations
 - Pearson Chemistry Chapter 8 Covalent Bonding Answers User Reviews and Ratings
 - Pearson Chemistry Chapter 8 Covalent Bonding Answers and Bestseller Lists
5. Accessing Pearson Chemistry Chapter 8 Covalent Bonding Answers Free and Paid eBooks
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 - Pearson Chemistry Chapter 8 Covalent Bonding Answers eBook Subscription Services
 - Pearson Chemistry Chapter 8 Covalent Bonding Answers Budget-Friendly Options
6. Navigating Pearson Chemistry Chapter 8 Covalent Bonding Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Chapter 8 Covalent Bonding Answers Compatibility with Devices
 - Pearson Chemistry Chapter 8 Covalent Bonding Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Highlighting and Note-Taking Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Interactive Elements Pearson Chemistry Chapter 8 Covalent Bonding Answers

8. Staying Engaged with Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Chapter 8 Covalent Bonding Answers
9. Balancing eBooks and Physical Books Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Chapter 8 Covalent Bonding Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Setting Reading Goals Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Fact-Checking eBook Content of Pearson Chemistry Chapter 8 Covalent Bonding Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
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