Name: Key

The number of moles of substances, when compared and reduced to the lowest terms. yields a mole ratio, as shown in the sample below. Indicate the mole ratio for each of the following equations:

$$4H_2 + 2O_2 \Rightarrow 4H_2O$$
 2: 1: 2

 $H_2 + S \Rightarrow H_2S$
 1: 1: 1

 $4AI + 3O_2 \Rightarrow 2 AI_2O_3$
 4: 3: 2

 $3V_1 + 6I_1CI \Rightarrow 3V_1CI_2 + 3H_2$
 1: 2: 1: 1

 $SO_2 + \frac{1}{2}O_2 \Rightarrow SO_3$
 2: 1: 2

 $6H_2 + 3C \Rightarrow 3CH_4$
 2: 1: 1

FW= 17.00/mol

Convert the following masses to moles by using the information on your periodic table.

a)
$$34.0 \text{ g NH}_3$$
 FW = 17.05 feet
 $34.0 \text{ g (} \frac{\text{I mol}}{17.0 \text{ g}} \text{)} = 2.00 \text{ mol}$

b) $51.4 \text{ g Na}_2\text{O}$ FW = 62.09 fmol

 $51.4 \text{ g (} \frac{\text{I mol}}{62.0 \text{ g}} \text{)} = 0.829 \text{ mol}$

c) $4.50 \text{ g II}_2\text{SO}_4$ FW = 98.1 g fmol

 $4.50 \text{ g (} \frac{\text{I mol}}{98.1 \text{ g}} \text{)} = 0.0459 \text{ mol}$

d) $229 \text{ g Ca(NO}_3)_2$ FW = 164.19 fmol

 $229 \text{ g (} \frac{\text{I mol}}{164.19} \text{)} = 1.40 \text{ mol}$

c) $79.5 \text{ g NH}_4\text{OH}$ FW = 35.0 g fmol
 $79.5 \text{ g (} \frac{\text{I mol}}{35.0 \text{ g (}} \text{)} = 2.27 \text{ mol}$

Convert the following to grams.

34.0 g NH1

a)
$$17.1 \text{ moles } H_2O = 18.09 \text{ mol}$$
 308 g .

 $17.1 \text{ mol} \times \left(\frac{18.09}{1 \text{ mol}}\right) = 3089$

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Mike Jess

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