

Chemistry Student Edition - Basic Answer Key

Chapter 17: Thermochemistry

Heat Flow

Questions

1. What is true of all chemical processes?
2. How can you distinguish a system and its surroundings?
3. What are the required characteristics of an isolated system?
4. Explain the difference between exothermic and endothermic processes.
5. Indicate which of the following processes are exothermic and which are endothermic:
 - a. ice melting
 - b. burning a candle
 - c. cooking an egg
 - d. the evaporation of sweat
6. Calculate the amount of heat needed to increase the temperature of 125 g of water from 22°C to 59°C.
7. Calculate the specific heat of copper, given that 204.75 J of energy raises the temperature of 15 g of copper from 35°C to 70°C.
8. 432 J of energy is required to raise the temperature of a block of aluminum from 20°C to 60°C. Calculate the mass of aluminum present.
9. 300 g of liquid ethanol at 22°C is supplied with 6480 J of heat. What is the final temperature of the ethanol? (The specific heat of ethanol is 2.44 J/g·°C).

Answers

1. All chemical processes involve the transfer of heat to or from the system.
2. The system includes the components involved in a reaction, while the surroundings include everything else.
3. An isolated system can exchange neither heat or matter with its surroundings.
4. Exothermic processes release heat to the surroundings, while endothermic processes require heat from the surroundings.
5.
 - a. Endothermic
 - b. Exothermic
 - c. Endothermic
 - d. Endothermic
6. mass = 125 g
 $\Delta T = 37^{\circ}\text{C}$
specific heat of water = 4.18 J/g $^{\circ}\text{C}$
$$\Delta H = 125\text{g} \times \frac{4.18\text{J}}{\text{g}^{\circ}\text{C}} \times 37^{\circ}\text{C}$$
$$= 19332.5 \text{ J or } 19.3 \text{ kJ}$$

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