

$$2. [\text{OH}^-] = 1.0 \times 10^{-14} \text{ M}$$

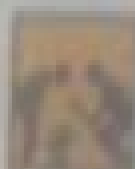
$$[\text{H}^+] = ?$$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$$1.0 \times 10^{-14} = [\text{H}^+](1.0 \times 10^{-14})$$

$$[\text{H}^+] = 1.0 \times 10^{-14} \text{ M}$$

Because $1.0 \times 10^{-14} \text{ M} < 1.0 \times 10^{-7} \text{ M}$, the solution is acidic.



Self Assessment Exercise 10.3

1. A soft drink has $[\text{H}^+] = 3 \times 10^{-4} \text{ M}$. Is the drink acidic, basic, or neutral?
2. Ordinary vinegar is approximately 1M CH_3COOH . Concentration of H^+ is $4.2 \times 10^{-3} \text{ M}$. Is vinegar acidic, basic, or neutral?
3. A student measures the $[\text{OH}^-]$ of a solution and obtains a value of $1.0 \times 10^{-10} \text{ M}$. Is the solution acidic, basic, or neutral?

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Rachel S Tattersall



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 - The Rise of Digital Reading Pearson Chem Tutor 10 Self Check Activity
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chem Tutor 10 Self Check Activity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chem Tutor 10 Self Check Activity
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chem Tutor 10 Self Check Activity
 - Personalized Recommendations
 - Pearson Chem Tutor 10 Self Check Activity User Reviews and Ratings
 - Pearson Chem Tutor 10 Self Check Activity and Bestseller Lists
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 - Pearson Chem Tutor 10 Self Check Activity Enhanced eBook Features
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 - Highlighting and Note-Taking Pearson Chem Tutor 10 Self Check Activity
 - Interactive Elements Pearson Chem Tutor 10 Self Check Activity
- 8. Staying Engaged with Pearson Chem Tutor 10 Self Check Activity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chem Tutor 10 Self Check Activity
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chem Tutor 10 Self Check Activity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chem Tutor 10 Self Check Activity
 - Setting Reading Goals Pearson Chem Tutor 10 Self Check Activity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chem Tutor 10 Self Check Activity
 - Fact-Checking eBook Content of Pearson Chem Tutor 10 Self Check Activity
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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