

AQA Chemistry (8462) from 2016 Topics C4.1 Atomic structure and the periodic table				
Topic	Student Checklist	R	A	G
4.1.1 A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes	State that everything is made of atoms and recall what they are			
	Describe what elements and compounds are			
	State that elements and compounds are represented by symbols; and use chemical symbols and formulae to represent elements and compounds			
	Write word equations and balanced symbol equations for chemical reactions, including using appropriate state symbols			
	HT ONLY: Write balanced half equations and ionic equations			
	Describe what a mixture is			
	Name and describe the physical processes used to separate mixtures and suggest suitable separation techniques			
	Describe how the atomic model has changed over time due to new experimental evidence, inc discovery of the atom and scattering experiments (inc the work of James Chadwick)			
	Describe the difference between the plum pudding model of the atom and the nuclear model of the atom			
	State the relative charge of protons, neutrons and electrons and describe the overall charge of an atom			
	State the relative masses of protons, neutrons and electrons and describe the distribution of mass in an atom			
	Calculate the number of protons, neutrons and electrons in an atom when given its atomic number and mass number			
	Describe isotopes as atoms of the same element with different numbers of neutrons			
	Define the term relative atomic mass and why it takes into account the abundance of isotopes of the element			
	Calculate the relative atomic mass of an element given the percentage abundance of its isotopes			
4.1.2 The periodic table	Describe how electrons fill energy levels in atoms, and represent the electron structure of elements using diagrams and numbers			
	Recall how the elements in the periodic table are arranged			
	Describe how elements with similar properties are placed in the periodic table			
	Explain why elements in the same group have similar properties and how to use the periodic table to predict the reactivity of elements			
	Describe the early attempts to classify elements			
	Explain the creation and attributes of Mendeleev's periodic table			
	Identify metals and non-metals on the periodic table, compare and contrast their properties			
	Explain how the atomic structure of metals and non-metals relates to their position in the periodic table			
	Describe noble gases (group 0) and explain their lack of reactivity			
	Describe the properties of noble gases, including boiling points, predict trends down the group and describe how their properties depend on the outer shell of electrons			
	Describe the reactivity and properties of group 1 alkali metals with reference to their electron arrangement and predict their reactions			
	Describe the properties of group 7 halogens and how their properties relate to their electron arrangement, including trends in molecular mass, melting and boiling points and reactivity			
	Describe the reactions of group 7 halogens with metals and non-metals			
	Chem ONLY: Describe the properties of transition metals and compare them with group 1 elements, including melting points and densities, strength and hardness, and reactivity (for Cr, Mn Fe, Co, Ni & Cu)			

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Table of Contents Pixl Club Paper 2014 Aqa Style

1. Understanding the eBook Pixl Club Paper 2014 Aqa Style
 - The Rise of Digital Reading Pixl Club Paper 2014 Aqa Style
 - Advantages of eBooks Over Traditional Books
2. Identifying Pixl Club Paper 2014 Aqa Style
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pixl Club Paper 2014 Aqa Style
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pixl Club Paper 2014 Aqa Style
 - Personalized Recommendations
 - Pixl Club Paper 2014 Aqa Style User Reviews and Ratings
 - Pixl Club Paper 2014 Aqa Style and Bestseller Lists
5. Accessing Pixl Club Paper 2014 Aqa Style Free and Paid eBooks

- Pixl Club Paper 2014 Aqa Style Public Domain eBooks
- Pixl Club Paper 2014 Aqa Style eBook Subscription Services
- Pixl Club Paper 2014 Aqa Style Budget-Friendly Options
- 6. Navigating Pixl Club Paper 2014 Aqa Style eBook Formats
 - ePub, PDF, MOBI, and More
 - Pixl Club Paper 2014 Aqa Style Compatibility with Devices
 - Pixl Club Paper 2014 Aqa Style Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pixl Club Paper 2014 Aqa Style
 - Highlighting and Note-Taking Pixl Club Paper 2014 Aqa Style
 - Interactive Elements Pixl Club Paper 2014 Aqa Style
- 8. Staying Engaged with Pixl Club Paper 2014 Aqa Style
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pixl Club Paper 2014 Aqa Style
- 9. Balancing eBooks and Physical Books Pixl Club Paper 2014 Aqa Style
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pixl Club Paper 2014 Aqa Style
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pixl Club Paper 2014 Aqa Style
 - Setting Reading Goals Pixl Club Paper 2014 Aqa Style
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pixl Club Paper 2014 Aqa Style
 - Fact-Checking eBook Content of Pixl Club Paper 2014 Aqa Style
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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