

## Polyatomic Ions Worksheet

### Polyatomic Practice

1. Name or write the formula for the following polyatomic ions

|              |       |                     |       |
|--------------|-------|---------------------|-------|
| sulfate      | _____ | $\text{CO}_3^{2-}$  | _____ |
| nitrite      | _____ | $\text{MnO}_4^{1-}$ | _____ |
| perphosphate | _____ | $\text{SO}_3^{2-}$  | _____ |
| hypiodite    | _____ | $\text{BrO}_2^{1-}$ | _____ |
| chlorite     | _____ | $\text{CO}_3^{2-}$  | _____ |
| phosphite    | _____ | $\text{PO}_3^{2-}$  | _____ |
| percarbonate | _____ | $\text{ClO}^{1-}$   | _____ |
| bromate      | _____ | $\text{IO}_3^{1-}$  | _____ |
| hyposulfite  | _____ | $\text{PO}_4^{2-}$  | _____ |
| permanganate | _____ | $\text{NO}_3^{1-}$  | _____ |
| carbonate    | _____ | $\text{SO}_4^{2-}$  | _____ |

2. Name or write the formula for the following Type I polyatomic ionic compounds

|                         |       |                                               |       |
|-------------------------|-------|-----------------------------------------------|-------|
| beryllium hydroxide     | _____ | $\text{Ba}(\text{IO}_3)_2$                    | _____ |
| sodium nitrite          | _____ | $\text{Ga}(\text{CNO})_3$                     | _____ |
| ammonium chloride       | _____ | $\text{Ag}_2\text{SO}_4$                      | _____ |
| calcium bisulfate       | _____ | $\text{MgCO}$                                 | _____ |
| rubidium perchlorate    | _____ | $\text{NH}_2\text{NO}_2$                      | _____ |
| strontium sulfite       | _____ | $\text{Al}(\text{C}_2\text{H}_3\text{O}_5)_3$ | _____ |
| aluminum acetate        | _____ | $\text{SrSO}_4$                               | _____ |
| ammonium nitrate        | _____ | $\text{RbClO}_2$                              | _____ |
| magnesium hypocarbonite | _____ | $\text{Ca}(\text{HSO}_4)_2$                   | _____ |
| silver hyposulfite      | _____ | $\text{NH}_4\text{Cl}$                        | _____ |
| gallium cyanate         | _____ | $\text{NaNO}$                                 | _____ |
| barium iodate           | _____ | $\text{Be}(\text{OH})_2$                      | _____ |

# Polyatomic Ions Extension Questions Answers

**RS Peters**



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**Chemistry** Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order *Surface and Defect Properties of Solids* M W Roberts, John M Thomas, 2007-10-31 Specialist Periodical Reports provide systematic and detailed review coverage of progress in the major areas of chemical research Written by experts in their specialist fields the series creates a unique service for the active research chemist supplying regular critical in depth accounts of progress in particular areas of chemistry For over 90 years The Royal Society of Chemistry and its predecessor the Chemical Society have been publishing reports charting developments in chemistry which originally took the form of Annual Reports However by 1967 the whole spectrum of chemistry could no longer be contained within one volume and the series Specialist Periodical Reports was born The Annual Reports themselves still existed but were divided into two and subsequently three volumes covering Inorganic Organic and Physical Chemistry For more general coverage of the highlights in chemistry they remain a must Since that time the SPR series has altered according to the fluctuating degree of activity in various fields of chemistry Some titles have remained unchanged while others have altered their emphasis along with their titles some have been combined under a new name whereas others have had to be discontinued The current list of Specialist Periodical Reports can be seen on the inside flap of this volume *Chemistry* Trace Jordan, Neville R. Kallenbach, 2017 Chemistry The Molecules of Life offers chemical insights within the context of health pharmaceuticals and the function of biological molecules The contextualized presentation of topics gives students a broad introduction to chemistry and helps them to see the relevance of chemistry to their personal lives **State Selected and State-to-State Ion-Molecule Reaction Dynamics, Volume 82, Part 2** Michael Baer, Cheuk-Yiu Ng, 1992-04-16 The Advances in Chemical Physics series provides the chemical physics and physical chemistry fields with a forum for critical authoritative evaluations of advances in every area of the discipline Filled with cutting edge research reported in a cohesive manner not found elsewhere in the literature each volume of the Advances in

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Multiphoton Ionization in Polyatomic Molecules David Hubert Parker, 1979 *Saline Water Conversion Report* United States. Office of Saline Water, 1968 *Electrons in Finite and Infinite Structures* P. Phariseau, 2013-06-29 This book contains the transcripts of the lectures presented at the NATO Advanced Study Institute on Electrons in Finite and Infinite Structures held at the State University of Ghent Belgium August 30 September 11 1976 Over the last few years substantial progress has been made in the description and the understanding of the behavior of electrons in extended bodies This includes the study of the energy spectrum of electrons in large molecules perfect as well as imperfect crystals and disordered alloys Not only local potential techniques but also the many body aspects are discussed in detail As atomic molecular and solid state physics involve common techniques and insights we believe that physicists and chemists active in these fields have benefited from these lectures and the interchange of ideas during the course The aim of the Institute was to familiarize young scientists in the field with the current state of the art and to indicate in which areas advances may be expected in the near future The ASI consisted of two parts detailed instructional and review lectures over the whole period and some evening sessions where

the participants were offered the opportunity to present their own work and discuss their ideas with senior scientists. Since the Institute took place a few weeks after Prof Dr John C Slater was suddenly taken from our scientific community it was a great honor for us to dedicate this course on behalf of the organizing committee to the late John C Slater.

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