

Name: Marcie Hall Period: 3 Date: 5/1/23



Note-Taking Guide (CER)

Activity Title: Aquatic Ecosystems- Developing

Activity Scoreboard when finished, circle highest score achieved					
Attempt #	1	2	3	4	5
Date					
Score					

Claim:
There are different aquatic systems with diverse organisms.

Evidence	Reasoning/Justification
Aquatic ecosystems have abiotic - temp, light, salinity, depth	Different organisms can tolerate different factors
Some organisms prefer fresh water vs. saltwater. Bull sharks + manatees tolerate both	Bull sharks and manatees are very different, but tolerate same conditions
Depth reduces light and increases pressure. Organisms deal with this differently.	Many types of adaptations exist to survive in these conditions
The geography of an aquatic environment can determine the life forms present.	Continental shelf gets lots of light, so there's many living things. Not so in the abyss.
Photosynthesis takes place in the neritic zone.	It is the most productive area of the ocean. Plankton are base of food chain.
Metabolism is affected by temp. Many species are cold-blooded.	Organisms can migrate or must have adaptations to survive.

Note Taking Guide 60answers

Christian G. Meyer

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 1300 \text{ kJ}$
 $2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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CO_2 g 2 H_2O l h 1411 kJ C_2H_6 g $3\frac{1}{2}$ O_2 g 2 CO_2 g 3 H_2O l h 1560 kJ

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O h 55.8 kJ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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