

- (d) Write the chemical equation of the reaction which takes place on reacting  $\text{XNO}_3$  solution and common salt solution giving the physical states of all the reactants and products.
- (e) What type of chemical reaction is illustrated by the above equation?
64. Two metals X and Y form the salts  $\text{XSO}_4$  and  $\text{Y}_2\text{SO}_4$  respectively. The solution of salt  $\text{XSO}_4$  is blue in colour whereas that of  $\text{Y}_2\text{SO}_4$  is colourless. When barium chloride solution is added to  $\text{XSO}_4$  solution, then a white precipitate Z is formed alongwith a salt which turns the solution green. And when barium chloride solution is added to  $\text{Y}_2\text{SO}_4$  solution, then the same white precipitate Z is formed alongwith colourless common salt solution.
- (a) What could the metals X and Y be?
- (b) Write the name and formula of salt  $\text{XSO}_4$ .
- (c) Write the name and formula of salt  $\text{Y}_2\text{SO}_4$ .
- (d) What is the name and formula of white precipitate Z?
- (e) Write the name and formula of the salt which turns the solution green in the first case.
65. A red-brown metal X forms a salt  $\text{XSO}_4$ . When hydrogen sulphide gas is passed through an aqueous solution of  $\text{XSO}_4$ , then a black precipitate of  $\text{XS}$  is formed alongwith sulphuric acid solution.
- (a) What could the salt  $\text{XSO}_4$  be?
- (b) What is the colour of salt  $\text{XSO}_4$ ?
- (c) Name the black precipitate  $\text{XS}$ .
- (d) By using the formula of the salt obtained in (a) above, write an equation of the reaction which takes place when hydrogen sulphide gas is passed through its aqueous solution.
- (e) What type of chemical reaction takes place in this case?
66. When a strip of red-brown metal X is placed in a colourless salt solution  $\text{YNO}_3$ , then metal Y is set free and a blue coloured salt solution  $\text{X(NO}_3)_2$  is formed. The liberated metal Y forms a shining white deposit on the strip of metal X.
- (a) What do you think metal X is?
- (b) Name the salt  $\text{YNO}_3$ .
- (c) What could be metal Y?
- (d) Name the salt  $\text{X(NO}_3)_2$ .
- (e) What type of reaction takes place between metal X and salt solution  $\text{YNO}_3$ ?
67. A metal salt  $\text{MX}$  when exposed to light splits up to form metal M and a gas  $\text{X}_2$ . Metal M is used in making ornaments whereas gas  $\text{X}_2$  is used in making bleaching powder. The salt  $\text{MX}$  is itself used in black and white photography.
- (a) What do you think metal M is?
- (b) What could be gas  $\text{X}_2$ ?
- (c) Name the metal salt  $\text{MX}$ .
- (d) Name any two salt solutions which on mixing together can produce a precipitate of salt  $\text{MX}$ .
- (e) What type of chemical reaction takes place when salt  $\text{MX}$  is exposed to light? Write the equation of the reaction.

### ANSWERS

1. Decomposition reaction      4.  $\text{Zn} + 2\text{AgNO}_3 \longrightarrow \text{Zn(NO}_3)_2 + 2\text{Ag}$       5. Rancidity      6. Anti-oxidants
11. Decomposition reactions (carried out by electricity)      14.  $2\text{AgNO}_3(\text{aq}) + \text{Cu}(\text{s}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$       15. (i) Decomposition      (ii) Combination      (iii) Decomposition      (iv) Decomposition      (v) Decomposition
- (c) Combination      16. (a) Combination      (b) Decomposition      (c) Decomposition      (d) Decomposition
- (e) Combination      17. (i) Displacement reaction      (ii) Combination reaction      (iii) Decomposition reaction
- (iv) Double displacement reaction      (v) Displacement reaction      18. (a) See page 29 (b) See page 29
19. (a) Displacement reaction      (b) Combination reaction      20. (a) Combination reaction
- (b) Displacement reaction      (c) Displacement reaction      (d) Decomposition reaction      (e) Double displacement reaction
21. (a)  $\text{H}_2\text{O}_2$       (b)  $\text{PbS}$       22.  $\text{H}_2\text{S}$       23. Substance oxidised :  $\text{H}_2\text{S}$  ; Substance reduced :  $\text{SO}_2$
24. (a) oxidation ; reduction      (b) reduction ; oxidation      (c) rancidity      25. (i) C      (ii)  $\text{ZnO}$       26. (b) Magnesium is oxidised ; Addition of oxygen to magnesium takes place      (c) (i)  $\text{HCl}$       (ii)  $\text{MnO}_2$       (iii)  $\text{MnO}_2$       (iv)  $\text{HCl}$
31.  $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$       (i) Zinc oxide      (ii) Carbon      34. (b)  $2\text{Cu} + \text{O}_2 \longrightarrow 2\text{CuO}$       (i) Cu      (ii)  $\text{O}_2$
35. (b)  $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$       (i) Mg      (ii)  $\text{O}_2$       37. (b)  $\text{Cu}(\text{s}) + 2\text{AgNO}_3(\text{aq}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$       39. (a)  $\text{H}_2$       (b)  $\text{CuO}$       (c)  $\text{CuO}$       (d)  $\text{H}_2$       43. (b) Rusting      (c) Oxidation      44. (b) Oxidation

# S Chand Lab Manual Class 5 Science

**S Nieto**



**S Chand Lab Manual Class 5 Science:**

The book delves into S Chand Lab Manual Class 5 Science. S Chand Lab Manual Class 5 Science is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into S Chand Lab Manual Class 5 Science, encompassing both the fundamentals and more intricate discussions.

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    - Chapter 2: Essential Elements of S Chand Lab Manual Class 5 Science
    - Chapter 3: S Chand Lab Manual Class 5 Science in Everyday Life
    - Chapter 4: S Chand Lab Manual Class 5 Science in Specific Contexts
    - Chapter 5: Conclusion
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