

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
Mg(OH) ₂	<u>base</u>	H ₃ PO ₄	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO ₃	<u>salt</u>	Al(OH) ₃	<u>base</u>
HFO ₄	<u>acid</u>	KC ₂ H ₃ O ₂	<u>salt</u>
Ba(OH) ₂	<u>base</u>	NaCl	<u>salt</u>

Note Taking Worksheet Acids Bases And Salts

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Reviewing **Note Taking Worksheet Acids Bases And Salts**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Note Taking Worksheet Acids Bases And Salts**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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