

6.2.1 EXERCISES

In Exercises 1 - 15, expand the given logarithm and simplify. Assume when necessary that all quantities represent positive real numbers.

1. $\ln(x^3y^2)$

2. $\log_2 \left(\frac{128}{x^2 + 4} \right)$

3. $\log_5 \left(\frac{z}{25} \right)^3$

4. $\log(1.23 \times 10^{37})$

5. $\ln \left(\frac{\sqrt{z}}{xy} \right)$

6. $\log_5 (x^2 - 25)$

7. $\log_{\sqrt{2}} (4x^3)$

8. $\log_{\frac{1}{3}} (9x(y^3 - 8))$

9. $\log (1000x^3y^5)$

10. $\log_3 \left(\frac{x^2}{81y^4} \right)$

11. $\ln \left(\sqrt[4]{\frac{xy}{ez}} \right)$

12. $\log_6 \left(\frac{216}{x^3y} \right)^4$

13. $\log \left(\frac{100x\sqrt{y}}{\sqrt[3]{10}} \right)$

14. $\log_{\frac{1}{2}} \left(\frac{4\sqrt[3]{x^2}}{y\sqrt{z}} \right)$

15. $\ln \left(\frac{\sqrt[3]{x}}{10\sqrt{yz}} \right)$

In Exercises 16 - 29, use the properties of logarithms to write the expression as a single logarithm.

16. $4\ln(x) + 2\ln(y)$

17. $\log_2(x) + \log_2(y) - \log_2(z)$

18. $\log_3(x) - 2\log_3(y)$

19. $\frac{1}{2}\log_3(x) - 2\log_3(y) - \log_3(z)$

20. $2\ln(x) - 3\ln(y) - 4\ln(z)$

21. $\log(x) - \frac{1}{3}\log(z) + \frac{1}{2}\log(y)$

22. $-\frac{1}{3}\ln(x) - \frac{1}{3}\ln(y) + \frac{1}{3}\ln(z)$

23. $\log_5(x) - 3$

24. $3 - \log(x)$

25. $\log_7(x) + \log_7(x - 3) - 2$

26. $\ln(x) + \frac{1}{2}$

27. $\log_2(x) + \log_4(x)$

28. $\log_2(x) + \log_4(x - 1)$

29. $\log_2(x) + \log_{\frac{1}{2}}(x - 1)$

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web in molecular compounds the atom binds each other through covalent bonds in salts it is held together with ionic bonds
these are the two types of bonds out of which every compound is made of 2 example of compounds example of compounds
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web molecular compounds many compounds do not contain ions but instead consist solely of discrete neutral molecules these
molecular compounds covalent compounds result when atoms share rather than transfer gain or lose electrons covalent
bonding is an important and extensive concept in chemistry and it will be treated in considerable

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it is ionic write the symbols for the ions involved a NF_3 b BaO c NH_4Cl d SrH_2 e PO_4^{3-} f Na_2O answer a answer b
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and nothing else so it is pure when we mix two different pure

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with the suffix ide the prefixes are as follows mono 1 di 2 tri 3 tetra 4

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substances as water H_2O H_2O and carbon dioxide CO_2 CO_2 these compounds are very different from ionic compounds like
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web may 28 2020 what is the total mass of hydrogen in each of the molecules a CH_4 b CHCl_3 c $\text{C}_{12}\text{H}_{10}\text{O}_6$ d $\text{CH}_3\text{CH}_2\text{CH}_2$

ch 2 ch 3 answer a answer b answer c answer d problem 4 4 1 3 calculate the molecular or formula mass of each of the following a p 4 b h 2 o c ca no 3 2 d ch 3 co 2 h acetic acid

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