

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES
(CHEMISTRY)**

TABLE 1 PHYSICAL CONSTANTS

NAME	SYMBOL	VALUE
Magnitude of charge on electron	e	$1,6 \times 10^{-19} \text{ C}$
Mass of an electron	m_e	$9,1 \times 10^{-31} \text{ kg}$
Standard pressure	p°	$1,01 \times 10^5 \text{ Pa}$
Molar gas volume at STP	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	T°	273 K
Avogadro's constant	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	F	$96\,500 \text{ C} \cdot \text{mol}^{-1}$

TABLE 2 CHEMISTRY FORMULAE

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$	$n = \frac{V}{V_m}$
$c = \frac{n}{V}$ OR $c = \frac{m}{MV}$	$K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at 25°C (298 K)	
$q = It$ $q = nF$	$E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$ $E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$	

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Rachel S Tattersall



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TABLE OF CONTENTS Solved Paper 2019 Dec Solved Paper 2019 July Solved Paper 2018 Dec Solved Paper 2016 Sept Solved Paper 2016 Feb Practice Sets 1 15 15 Practice Sets CTET Paper-2 Paper 2 Mas & Science Teacher Selection for

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TABLE OF CONTENTS
 Solved Paper 2019 Dec Solved Paper 2019 July Solved Paper 2018 Dec Solved Paper 2016 Sept Solved Paper 2016 Feb
 Practice Sets 1 15

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 Practice sets 1 15

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TABLE OF CONTENT

Solved Paper 2019 December Solved Paper 2019 July Solved Paper 2018 December Solved Paper 2016 September

Child Development and Pedagogy English Language and Pedagogy Hindi Bhasha evm Shiksha Shastra Mathematics and Pedagogy Science and Pedagogy Practice Sets 1 2

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1. Understanding the eBook Physical Science P2 Grade1exampler
 - The Rise of Digital Reading Physical Science P2 Grade1exampler
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science P2 Grade1exampler
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science P2 Grade1exampler
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science P2 Grade1exampler
 - Personalized Recommendations
 - Physical Science P2 Grade1exampler User Reviews and Ratings

- Physical Science P2 Grade1exampler and Bestseller Lists
- 5. Accessing Physical Science P2 Grade1exampler Free and Paid eBooks
 - Physical Science P2 Grade1exampler Public Domain eBooks
 - Physical Science P2 Grade1exampler eBook Subscription Services
 - Physical Science P2 Grade1exampler Budget-Friendly Options
- 6. Navigating Physical Science P2 Grade1exampler eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science P2 Grade1exampler Compatibility with Devices
 - Physical Science P2 Grade1exampler Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science P2 Grade1exampler
 - Highlighting and Note-Taking Physical Science P2 Grade1exampler
 - Interactive Elements Physical Science P2 Grade1exampler
- 8. Staying Engaged with Physical Science P2 Grade1exampler
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science P2 Grade1exampler
- 9. Balancing eBooks and Physical Books Physical Science P2 Grade1exampler
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science P2 Grade1exampler
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science P2 Grade1exampler
 - Setting Reading Goals Physical Science P2 Grade1exampler
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science P2 Grade1exampler
 - Fact-Checking eBook Content of Physical Science P2 Grade1exampler
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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