



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use permanent marker for writing and drawing.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

- 3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B or of A, B and C or of others as well? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (2)
- 3.1.4 What kind of the compound A or B is 1-hexanol? Explain based on physical trends? (2)
- 3.1.5 What compound B or C will have the higher vapour pressure? (2)
- 3.1.6 What compound B or C will have the higher surface tension? (2)

Compound C (M_r 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound C is 16,2 °C? (2)

- 3.2 Add table 3.2 to obtain a more complete picture of intermolecular forces and their relative contribution to the boiling point.

Compound	Boiling point (°C)	Relative contribution to boiling point
A	16,2	
B	21	

- 3.2.1 Is there any difference in the boiling point of A and B? (2)
- 3.2.2 Explain the reason for the difference in the boiling point of A and B. (2)

Physical Science Paper 2 Intermolecular Forces Investigation

Royal Society (Great Britain)



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