

Chapter 5 Study Guide- Physical Science
Test Monday, April 28th, 2014

1. What is a chemical bond? **force of attraction that holds elements together in order to form a unit**
2. Define chemical structure. **the arrangement of atoms in a substance.**
3. What is an ionic bond? **a chemical bond in which one atom loses an electron to form a positive ion and the other atom gains an electron to form a negative ion**
4. An ionic bond is between a **METAL** and a **NONMETAL** .
5. What is a covalent bond? **a chemical bond that involves sharing a pair of electrons between atoms in a molecule**
6. A covalent bond is between a **nonmetal** and a **nonmetal** .
7. What is a metallic bond? **a bond formed by the attraction between positively charged metal ions and the electrons around them.**
8. What are the polyatomic ions? **an ion made of two or more atoms.**
9. What is an ION? **a charged atom**
10. What is a CATION? **A positively charged atom**
11. What is an ANION? **A negatively charged atom**
12. Which part of an ionic compound should be written first? **Metal**

Name the following ionic compounds:

13. FeI_2 - **Iron Iodide**
14. MnF_3 - **Manganese Fluoride**
15. CrCl_2 - **Chromium Chloride**
16. CuS - **Copper Sulfide**

Name the following covalent compounds:

17. As_2O_5 - **diarsenic pentaoxide**
18. SiI_4 - **silicon tetraiodide**
19. P_4S_3 - **tertraphosphorus trisulfide**
20. P_4O_{10} - **tetraphosphorus decoxide**

Write the equation for the following ionic compounds:

21. Lithium Oxide- **Li_2O**
22. Magnesium Nitride- **Mg_3N_2**
23. Sodium Phosphide- **Na_3P**
24. Calcium Bromide - **CaBr_2**

Define/Explain the following

25. What is a subscript? **number written below, tells you how many of each element is present in compound/molecule**
26. What is a superscript? **number written above the element symbol, this is how the oxidation number (charge) is displayed.**

Physical Science Chemical Bond Study Guide

George DeBoer



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Table of Contents **Physical Science Chemical Bond Study Guide**

1. Understanding the eBook **Physical Science Chemical Bond Study Guide**
 - The Rise of Digital Reading **Physical Science Chemical Bond Study Guide**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Physical Science Chemical Bond Study Guide**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Physical Science Chemical Bond Study Guide**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Physical Science Chemical Bond Study Guide**
 - Personalized Recommendations
 - **Physical Science Chemical Bond Study Guide** User Reviews and Ratings
 - **Physical Science Chemical Bond Study Guide** and Bestseller Lists

5. Accessing Physical Science Chemical Bond Study Guide Free and Paid eBooks
 - Physical Science Chemical Bond Study Guide Public Domain eBooks
 - Physical Science Chemical Bond Study Guide eBook Subscription Services
 - Physical Science Chemical Bond Study Guide Budget-Friendly Options
6. Navigating Physical Science Chemical Bond Study Guide eBook Formats
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 - Physical Science Chemical Bond Study Guide Compatibility with Devices
 - Physical Science Chemical Bond Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chemical Bond Study Guide
 - Highlighting and Note-Taking Physical Science Chemical Bond Study Guide
 - Interactive Elements Physical Science Chemical Bond Study Guide
8. Staying Engaged with Physical Science Chemical Bond Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chemical Bond Study Guide
9. Balancing eBooks and Physical Books Physical Science Chemical Bond Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chemical Bond Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chemical Bond Study Guide
 - Setting Reading Goals Physical Science Chemical Bond Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chemical Bond Study Guide
 - Fact-Checking eBook Content of Physical Science Chemical Bond Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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