

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

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X-kit Exam 2004 Physical Science ,2005 **The Chemical News and Journal of Physical Science** ,1900

Chemical News and Journal of Physical Science ,1877 **Master the GED: The Science Test**

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the physical psychological and social needs of children and youth based on their ages and maturational levels This groundbreaking resource covers a breadth of topics including bone development burnout gender and racial stereotypes injuries motor behavior and parental pressures Written by Robin S Vealey and Melissa A Chase the 16 chapters of Best Practice for Youth Sport are divided into four parts Part I Youth Sport Basics provides readers with the fundamental knowledge and background related to the history evolution and organization of youth sport Part II Maturation and Readiness for Youth Sport Participants is the core of understanding how and why youth sport is different from adult sport This part details why it is important to know when youth are ready to learn and compete Part III Intensity of Participation in Youth Sport examines the appropriateness of physical and psychological intensity at various developmental stages and the potential ramifications of overtraining overspecialization overstress and overuse The text concludes with part IV Social Considerations in Youth Sport which examines how youth sport coaches and parents can help create a supportive social environment so that children can maximize the enjoyment and benefits from youth sport In addition to 14 appendixes activities glossaries study questions and other resources that appear in Best Practice for Youth Sport the textbook is enhanced with instructor ancillaries a test package image bank and instructor guide that features a syllabus additional study questions and learning activities tips on teaching difficult concepts and additional readings and resources These specialized resources ensure that instructors will be ready for each class session with engaging materials Ancillaries are free to course adopters and available at www.HumanKinetics.com/BestPracticeForYouthSport Best Practice for Youth Sport provides readers with knowledge of sport science concerning youth sport and engages them through the use of anecdotes activities case studies and practical strategies Armed with the knowledge from this text students coaches parents administrators and others will be able to become active agents of social change in structuring and enhancing youth sport programs to meet the unique developmental needs of children making the programs athlete centered rather than adult centered so that they truly are all about the kids

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