

OCR A LEVEL CHEMISTRY DEFINITIONS

Module 2:	
Term	Definition
Isotope	atoms of the same element with different numbers of neutrons and different masses
Relative isotopic mass	Mass compared with 1/12th mass of carbon-12
Relative atomic mass	weighted mean mass compared with 1/12th mass of carbon-12 atom.
Empirical formula	the simplest whole number ratio of atoms of each element present in a compound
Molecular formula	the number and type of atoms of each element in a molecule
Acid/Bronsted Lowry Acid	release H ⁺ ions in aqueous solution/Proton donor
Alkalis/ Bronsted Lowry Base	release OH ⁻ ions in aqueous solution/proton acceptor
Strong acid	Fully dissociates to release H ⁺ ions in aqueous solution
Weak acid	Partially dissociates to release H ⁺ ions in aqueous solution
Orbital	a region around the nucleus that can hold up to two electrons, with opposite spins
Ionic Bond	electrostatic attraction between positive and negative ions
Covalent Bond	strong electrostatic attraction between a shared pair of electrons and the nuclei of the bonded atoms
Metallic Bond	strong electrostatic attraction between cations (positive ions) and delocalised electrons
Electronegativity	ability of an atom to attract the bonding electrons in a covalent bond
Module 3:	
First Ionisation Energy	The energy required to remove 1 mol of electrons from 1 mol of gaseous atoms to form 1 mol of gaseous 1+ ions
Second Ionisation Energy	The energy required to remove 1 mol of electrons from 1 mol of gaseous 1+ ions to form 1 mol of gaseous 2+ ions
Disproportionation	oxidation and reduction of the same element
Standard conditions	298K, 1atm/100kPa, 1 moldm ⁻³ solutions
Standard enthalpy of formation	the enthalpy change when 1 mole of a compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion	the enthalpy change when 1 mole of a substance is completely burned in oxygen under standard conditions
Standard enthalpy of solution	the enthalpy change when 1 mole of an ionic solid is dissolved in a large volume of water under standard conditions
Standard enthalpy of hydration	the enthalpy change when 1 mole of gaseous ions is dissolved in a large volume of water under standard conditions
Standard enthalpy of atomisation	the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state under standard conditions
Standard enthalpy of neutralisation	the enthalpy change when 1 mole of an acid reacts with 1 mole of an alkali under standard conditions
Standard enthalpy of precipitation	the enthalpy change when 1 mole of an ionic solid is formed from its ions in solution under standard conditions
Standard enthalpy of redox	the enthalpy change when 1 mole of a substance is oxidised or reduced under standard conditions
Standard enthalpy of formation of a compound	the enthalpy change when 1 mole of a compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion of a compound	the enthalpy change when 1 mole of a compound is completely burned in oxygen under standard conditions
Standard enthalpy of solution of a compound	the enthalpy change when 1 mole of a compound is dissolved in a large volume of water under standard conditions
Standard enthalpy of hydration of a compound	the enthalpy change when 1 mole of a compound is dissolved in a large volume of water under standard conditions
Standard enthalpy of atomisation of a compound	the enthalpy change when 1 mole of a compound is broken down into its constituent atoms under standard conditions
Standard enthalpy of neutralisation of an acid	the enthalpy change when 1 mole of an acid reacts with 1 mole of an alkali under standard conditions
Standard enthalpy of precipitation of an ionic solid	the enthalpy change when 1 mole of an ionic solid is formed from its ions in solution under standard conditions
Standard enthalpy of redox of a substance	the enthalpy change when 1 mole of a substance is oxidised or reduced under standard conditions

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AN Whitehead



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The Top Books of the Year Ocr Alevel F32past Papers 23 May The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous captivating novels enthralling the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the engaging narratives that have enthralled audiences this year. Ocr Alevel F32past Papers 23 May : Colleen Hoover's "It Ends with Us" This heartfelt tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover expertly weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can triumph. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This spellbinding historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids absorbing storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Discover the Magic : Delia Owens "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and suspenseful novel that will keep you guessing until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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