

2-2

Possible interactions are:

- (a) ~~Electrostatic~~ interactions (because of double bonds).
- (b) Electrostatic repulsion, attractive (because of CH_2- groups). However, because of interaction between CH_2- groups it is unlikely the difference should be great.

$$\text{One } \begin{array}{c} \text{H} \\ | \\ \text{C}=\text{O} \end{array} \text{ has } \text{C}-\text{O} = 0 \quad \text{One } \text{H} \text{ has } \text{C}-\text{O} = 0 \\ \text{C}=\text{O} \text{ has } \text{C}-\text{O} = 0 \quad \text{One } \text{H} \text{ has } \text{C}-\text{O} = 0 \end{array}$$

2-3

Graphically, we find from the equation using perturbation and values. Varying α from 0 to 100%, all the values



of the approximation are zero in successive until and \mathbf{B} comes into the same distance, then alignment turns out to be the same. Another situation is possible with methyls interacting so that the repeat distance would be doubled.

2-4

Refer to Table 2-1, Fig. 2-3.
Polyisobutylene has cyclohexane:

$$X = 0.136, \quad V_0 = 108, \quad \omega_0 = 0.10$$

$$N\omega = 2.45 \times 10^{10}, \quad \text{so } \omega_0 = 2.45 \times 10^{10} \text{ rad/sec}^2$$

From the reference:

$$\frac{X}{V_0} = 0.557, \quad V_0 = \frac{108}{X}, \quad N\omega = 2.45 \times 10^{10}, \quad \text{then } \omega_0 = 0.21$$

2-5. $\text{m} = 100\text{M}_0$, $\text{methyl/methyl} = \text{lightest/lightest link}$

$N\omega_0 = 100\text{L}_0\omega_0$, per repeat unit ≈ 30 (for $\text{CH}_2-\text{O}-\text{CH}_2$ which is 3 atoms)

$\omega_0/\omega_0 = 100\text{L}_0\omega_0$, therefore cross-links $\approx \frac{1}{3}$ (per 3 atoms)

$$= 5513 - \omega = 10^{10} \text{ rad/sec}^2$$

$$V = 102 \times 1.21 \times 10^{-10} \text{ sec}^2 = 1.23 \times 10^{-7}$$

From Fig. 2-3, $\omega_0 = 0.069$

Because addition polymer, methyl , $\text{Total volume} = 1 \text{ cm}^3$

$$1.0 \times 10^{-7} \times \omega = \omega = 10^{10} \text{ rad/sec}^2 = 0.069$$

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Lei Huang

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1. Understanding the eBook Rodriguez Solution Manual
 - The Rise of Digital Reading Rodriguez Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Rodriguez Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Rodriguez Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Rodriguez Solution Manual
 - Personalized Recommendations
 - Rodriguez Solution Manual User Reviews and Ratings
 - Rodriguez Solution Manual and Bestseller Lists
5. Accessing Rodriguez Solution Manual Free and Paid eBooks
 - Rodriguez Solution Manual Public Domain eBooks
 - Rodriguez Solution Manual eBook Subscription Services
 - Rodriguez Solution Manual Budget-Friendly Options
6. Navigating Rodriguez Solution Manual eBook Formats

- ePub, PDF, MOBI, and More
 - Rodriguez Solution Manual Compatibility with Devices
 - Rodriguez Solution Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Rodriguez Solution Manual
 - Highlighting and Note-Taking Rodriguez Solution Manual
 - Interactive Elements Rodriguez Solution Manual
8. Staying Engaged with Rodriguez Solution Manual
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rodriguez Solution Manual
9. Balancing eBooks and Physical Books Rodriguez Solution Manual
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rodriguez Solution Manual
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Rodriguez Solution Manual
- Setting Reading Goals Rodriguez Solution Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Rodriguez Solution Manual
- Fact-Checking eBook Content of Rodriguez Solution Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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