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# Rollei 500 Digital Cameras Owners Manual

**G Psacharopoulos**



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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^-$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{Fe}^{2+} + 5\text{H}_2\text{O} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 4\text{Fe}^{3+} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method  $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te} + 10\text{H}^+$   $11\text{PbO}_2 + \text{IO}_3^- + 2\text{H}^+ \rightarrow 12\text{H}^+ + \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h<sub>2</sub> o<sub>2</sub> h and e in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> 2h<sub>2</sub>no<sub>2</sub> c<sub>6</sub>h<sub>6</sub>o<sub>6</sub> 2h<sub>2</sub>o 2no choose 1 answer c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> and no<sub>2</sub> a c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> and no<sub>2</sub> no<sub>2</sub> and c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> b no<sub>2</sub> and c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> h and no<sub>2</sub> c h and

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