

Thermochemistry Chemistry Questions with Solutions

Q-1: Given that standard molar enthalpies of formation of NO(g) and $\text{NO}_2\text{(g)}$ are respectively 90.3 kJ/mol and 33.2 kJ/mol. Calculate the enthalpy change for the reaction $2\text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO}_2\text{(g)}$.

Answer: $\Delta H = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$

$$\Delta H = 2(33.2) - [2(90.3) + 0]$$

$$\Delta H = -114.2 \text{ kJ}$$

Note: Enthalpy of formation of free elements like O_2 is always 0.

Q-2: The enthalpy of formation of AgCl is obtained from the enthalpy change from which one of the following process?

- a) $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl(s)}$
- b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2\text{(g)} \rightarrow \text{AgCl(s)}$
- c) $\text{AgCl(s)} \rightarrow \text{Ag(s)} + \frac{1}{2} \text{Cl}_2\text{(g)}$
- d) $\text{Ag(s)} + \text{AuCl} \rightarrow \text{Au(s)} + \text{AgCl(s)}$

Answer: b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2\text{(g)} \rightarrow \text{AgCl(s)}$

Explanation: The enthalpy change accompanying the formation of one mole of a pure substance in its standard state from the constituents elements in their standard state is called enthalpy of formation. For Ag, the standard state is solid and for chlorine, it is gas. Thus, the process that gives the enthalpy of formation of AgCl is: $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2\text{(g)} \rightarrow \text{AgCl(s)}$

Q-3: ΔH and ΔE for the reaction, $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{Fe(s)} + 3\text{H}_2\text{O(l)}$ at constant temperature are related as

- a) $\Delta H = \Delta E$
- b) $\Delta H = \Delta E + RT$
- c) $\Delta H = \Delta E + 3RT$
- d) $\Delta H = \Delta E - 3RT$

Answer: d) $\Delta H = \Delta E - 3RT$

Explanation: For any chemical reaction,

$$\Delta H = \Delta E + \Delta_{\text{ng}}RT$$

Where, Δ_{ng} = total number of moles of gaseous product- total number of moles of gaseous reactants.

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Complete Preparation for the MCAT Aftab S. Hassan,James L. Flowers,1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress **Solutions Manual to Accompany Elements of**

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Draw Lewis structures for each of the following. 1. NO₃. 1-. 2. CH₄.

Skill Practice 26 Skill Practice 26. Name: Date: Hour: ____.

1. What does it mean to say that a bond is polar? One of the atoms ...

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What force of attraction does the second energy level of a phosphorus atom "feel" from the ...

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