

Stoichiometry questions (1)

Consider : $4\text{NH}_3 + 5\text{O}_2 \rightarrow 6\text{H}_2\text{O} + 4\text{NO}$

- How many moles of H_2O are produced if 0.176 mol of O_2 are used?

$$\# \text{ mol H}_2\text{O} = 0.176 \cancel{\text{ mol O}_2} \times \frac{6 \text{ mol H}_2\text{O}}{5 \cancel{\text{ mol O}_2}} = 0.2112 \text{ mol}$$

- How many moles of NO are produced in the reaction if 17 mol of H_2O are also produced?

$$\# \text{ mol NO} = 17 \cancel{\text{ mol H}_2\text{O}} \times \frac{4 \text{ mol NO}}{6 \cancel{\text{ mol H}_2\text{O}}} = 11.33 \text{ mol NO}$$

Notice that a correctly balanced equation is essential to get the right answer

Questions On Stoichiometry

Brendan G. Carr

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