Now electronic configuration Co3+:

(8 1 1 1 1 1			
3d 4		40	4.4

Thus, outer orbital or high spin complex of $[CoF_a]^{3-}$:

19211111	[32]	(4) (4) (4)	6 6		
34	4ls	4p		4.4	

(Six pairs of electrons from six fluoride ions.)

- (b) Dibromide bis (ethane-1, 2-diamine) Cobalt (III) is the IUPAC name of [CoBr₂(en)₂]* complex.
- (c) The number of ions produced from the complex:

[Co(NH₃)₆]Cl₂ in solution is three.

 $[Co(NH_3)_6]Cl_2 \longrightarrow \{Co(NH_3)_6\}^{+2} + 2 Cl^-$

7. (a) Differentiate between the following:*

$$[1\times 3=3]$$

- (i) Adsorption and Absorption
- (ii) Lyophobic Sol and Lyophilic Sol
- (iii) Multimolecular Colloid and Macromolecular colloid

OR

- (b) (l) Define the following terms:*
 - (i) Zeta Potential
 - (iii) Coagulation
 - (II) Why a negatively charged sol is obtained when AgNO₃ solution is added to KI solution?*
 [3]
- 8. Define transition metals. Why Zn, Cd and Hg are not called transition metals? How is the variability in oxidation states of transition metals different from that of p-block elements?
- Ans. Transition metals are the metals that have partially filled d-orbitals and having the ability to form stable cations.

Zn, Cd, Hg are not regarded as transition metals because they have completely filled d-subshell with outer electronic configuration (n-1) d^{10} ms^2 .

The variability in oxidation states of transition metals is due to the incomplete filling of d-orbitals in such a way, that their oxidation states differ from each other by unity (i.e., Fe²⁺ and Fe³⁺). While, in case of p-block elements, the oxidiation state differ by two units (i.e., +3 and +5).

- 9. (a) What happens when:
 - (i) Propanone is treated with CH₃MgBr and then hydrolysed?

- (ii) Ethanal is treated with excess ethanol and acid?
- (iii) Methanal undergoes Cannizzaro reaction?
- (b) Write the main product in the following reactions: [1 × 3 = 3]
- (i) 2 CH₃COC1 + (CH₄)₅Cd →
- (ii) CH-CH-CHO Zn(Hg)/Conc. HC)

Ans. (a) (i) When propanone is treated with CH₃MgBr followed by hydrolysis it gives 2-methylpropan-2-ol.

$$CH_3$$
 CH_3
 CH_3

(ii) When ethanal reacts with excess of ethanol in presence of trace amount of HCl then, 1, 1-Dithoxyethane will be formed.

(iii) When methanal undergoes Cannizzaro reaction then, formic acid and methyl alcohol are obtained.

HCHO Cook. NWOH + CH₃OH

(Methanal) (Formic acid) (Methanal) (Methanal) (Hornic acid) (Methanal) (Methanal)

(b) (i) 2 CH₃COCI + (CH₃)₂ Cd → O

2 CH₃ − C−CH₃ + CdCl₂

(I., 1-Diethoocwethane).

Thus, 2-Propanone is the main product.

Thus, propane is the main procleuct.

Thus, cyclohexane is the main product.

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Thomas Griffiths

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