

Directions: Using the Reference Tables for Chemistry, locate the following information.

1. Name C_5H_{12} . **Pentane** Table P + R
2. Write the equation for the decay of Kr-85. $^{85}\text{Kr} \rightarrow \beta^- + ^{85}\text{Rb}$ Table N
3. Explain how you know that NaPO_4 is soluble in water but NiCrO_4 is not. **Na group 1/ all chromates insoluble** Table F
4. What is the definition of STP, and give the values? **Standard Temperature (0°C or 273K) and Pressure (1 atm or 101.3 kPa)** Table A
5. Name, and give the formulas of the strongest and weakest bases. **NaOH / NH_3** Table L
6. Name $C_2H_3O_2^-$ or CH_3COO^- **acetate ion** Table E
7. What is the solubility of sulfur dioxide at 40 degrees Celsius, in grams? **~5g** Table G
8. What is the freezing point of fluorine? **53 K** Table S
9. What are the units for the heat of fusion, and what do they mean? **J/g heat lost per gram of substance** Table B
10. What is the symbol for the mole? **mol** Table D
11. What is the vapor pressure of water at 75°C? **38 kPa** Table H
12. How much heat does it take to convert 20g of water to steam at 100°C? **45200 J (45.2 kJ)** I
13. What is the prefix for 1/1000 of a meter? **Milli-** Table C
14. What is the molecular formula of ammonia? **NH_3** Table L
15. What is the formula for the permanganate ion? **MnO_4^-** Table E
16. Name CH_3COOH **ethanoic acid (acetic acid)** Table R
17. Write the symbol for a positron? **β^+** Table O
18. What is the half-life of Pu-239? **2.4×10^4 years** Table N
19. Is the formation of water, from its elements endothermic or exothermic? **exothermic** Table I
20. What is the atomic mass of silver? **108 g/mol** P-Table
21. How much heat is released when LiBr dissolves in water? **-48.83 kJ** Table I
22. Give the names and formulas of the strongest and weakest acids. **HCl / CH_3COOH** Table K
23. What is the general formula for alkynes? What does it mean? **C_nH_{2n-2}** Table Q

Reference Tables Scavenger Hunt Answers

Roman Wölfel



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Table of Contents Reference Tables Scavenger Hunt Answers

1. Understanding the eBook Reference Tables Scavenger Hunt Answers
 - The Rise of Digital Reading Reference Tables Scavenger Hunt Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reference Tables Scavenger Hunt Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Reference Tables Scavenger Hunt Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reference Tables Scavenger Hunt Answers
 - Personalized Recommendations
 - Reference Tables Scavenger Hunt Answers User Reviews and Ratings
 - Reference Tables Scavenger Hunt Answers and Bestseller Lists

5. Accessing Reference Tables Scavenger Hunt Answers Free and Paid eBooks
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 - Reference Tables Scavenger Hunt Answers eBook Subscription Services
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 - Reference Tables Scavenger Hunt Answers Compatibility with Devices
 - Reference Tables Scavenger Hunt Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reference Tables Scavenger Hunt Answers
 - Highlighting and Note-Taking Reference Tables Scavenger Hunt Answers
 - Interactive Elements Reference Tables Scavenger Hunt Answers
8. Staying Engaged with Reference Tables Scavenger Hunt Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reference Tables Scavenger Hunt Answers
9. Balancing eBooks and Physical Books Reference Tables Scavenger Hunt Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reference Tables Scavenger Hunt Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reference Tables Scavenger Hunt Answers
 - Setting Reading Goals Reference Tables Scavenger Hunt Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reference Tables Scavenger Hunt Answers
 - Fact-Checking eBook Content of Reference Tables Scavenger Hunt Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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