The de Broglie Wavelength

- * A particle of large mass and ordinary speed has too small a wavelength to be detected by conventional means.
- * A tiny particle—such as an electron—moving at typical speed has a detectable wavelength.

$$A = \frac{h}{p} = \frac{h}{mv} = \frac{6.63 \times 6163^{9} \times J10^{-34} \text{ J} \cdot \text{s}}{(911 \text{ kg}) \times (010^{9} \text{ kg}) \times (1 \times 10^{6} \text{ m/s})} = 7.29 \times 10^{-10} \text{ m}$$

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Robert M. Kacmarek, James K. Stoller, Al Heuer

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