

Oxidation-Reduction Titrations I: Determination of Oxalate

Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. If 0.4586 g of sodium oxalate, $\text{Na}_2\text{C}_2\text{O}_4$, requires 33.67 mL of a KMnO_4 solution to reach the end point, what is the molarity of the KMnO_4 solution? $0.4586\text{g} / 133.99\text{g/mol} = 0.003423$
2. Titration of an oxalate sample gave the following percentages: 15.33%, 15.55%, and 15.56%. Calculate the average and the standard deviation.
3. Why does the solution decolorize upon standing after the equivalence point has been reached?
4. Why is the KMnO_4 solution filtered, and why should it not be stored in a rubber-stoppered bottle?
5. What volume of 0.100 M KMnO_4 would be required to titrate 0.23 g of $\text{K}_2[\text{Cu}(\text{C}_2\text{O}_4)_2] \cdot 2\text{H}_2\text{O}$?
6. Calculate the percent $\text{C}_2\text{O}_4^{2-}$ in each of the following: $\text{H}_2\text{C}_2\text{O}_4$, $\text{Na}_2\text{C}_2\text{O}_4$, $\text{K}_2\text{C}_2\text{O}_4$, and $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3] \cdot 3\text{H}_2\text{O}$.

Redox Titration Pre Lab Questions Answers

A. Berka,J. Vulterin,J. Zýka



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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 Newer Redox Titrants A. Berka, J. Vulterin, J. Zýka, 2013-10-22 International Series of Monographs in Analytical Chemistry Volume 22 Newer Redox Titrants focuses on the processes reactions methodologies and approaches involved in the study of redox titrants The publication first offers information on potassium permanganate in alkaline solution and compounds of trivalent manganese including standard solutions indicator and review of determinations The text then ponders on compounds of trivalent copper and potassium hexacyanoferrate The book ponders on hypohalites hypochlorite and hypobromite chloramine T and bromine as well as standard solutions indicator and review of determinations The publication also takes a look at iodine monochloride periodic acid and its salts lead IV acetate compounds of pentavalent vanadium and iron III salts The compounds of trivalent cobalt hydrogen peroxide chromium II salts tin II chloride sodium arsenite and compounds of monovalent copper are also elaborated The publication is a reliable reference for readers interested in newer redox titrants

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