

BALANCING REDOX REACTIONS

STEP BY STEP IN TWO WAYS...



BOTH START THE SAME WAY

1. ASSIGN OXIDATION NUMBERS TO EACH ATOM



NOTE: OXIDATION NUMBERS SUM UP TO ZERO ALWAYS

ALWAYS HAVE AN OX TO GO AND ALWAYS HAVE TWO REDS

1. BALANCE THE OX TO GO & ALWAYS HAVE TWO REDS

2. WRITE YOUR EQUATION AS TWO HALF REACTIONS BALANCING ALL ATOMS EXCEPT H AND O

NOTE: ALWAYS OX. + A REDUCING AGENT AND RED. + A OXIDIZING AGENT



HALF-REACTION METHOD (in acidic conditions)

3. BALANCE O WITH H₂O



4. BALANCE H WITH H⁺



5. BALANCE CHARGES WITH ELECTRONS

NOTE: ALWAYS HAVE POSITIVE ON OXIDATION & ALWAYS HAVE NEGATIVE ON RED



6. MAKE SURE THE NUMBER OF ELECTRONS IS THE SAME

NOTE: YOU CAN BALANCE WITH AN ELECTRONS ON OXIDATION & ALWAYS HAVE POSITIVE ON RED



7. PUT BOTH REACTIONS BACK TOGETHER



8. CHECK FOR ANYTHING THAT IS BOTH A PRODUCT & A REACTANT

NOTE: ALWAYS BALANCE WITH ONE ATOM AND CHECK CHARGES



BUT WHAT ABOUT BASIC CONDITIONS?

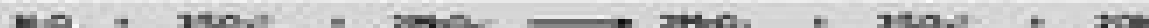
WHEN WE SAY "ACIDIC" OR "BASIC" WE ARE USING TWO APPROXIMATE DEFINITIONS:

ACIDIC CONDITIONS HAVE H⁺ AND BASIC CONDITIONS HAVE OH⁻

SO WHEN WE FIND OUR REACTION FROM ABOVE WITH A REACTION PERFORMED IN BASIC CONDITIONS:

8. ADD THE NUMBER OF OH⁻ TO THE NUMBER OF H⁺ IN EACH SIDE

NOTE: ALWAYS HAVE OH⁻ AND OH⁻ ON THE SAME SIDE THAT YOU WANT



Redox Reactions Teacher Guide And Answers

R Pring



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Table of Contents Redox Reactions Teacher Guide And Answers

1. Understanding the eBook Redox Reactions Teacher Guide And Answers
 - The Rise of Digital Reading Redox Reactions Teacher Guide And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox Reactions Teacher Guide And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox Reactions Teacher Guide And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Reactions Teacher Guide And Answers
 - Personalized Recommendations
 - Redox Reactions Teacher Guide And Answers User Reviews and Ratings
 - Redox Reactions Teacher Guide And Answers and Bestseller Lists
5. Accessing Redox Reactions Teacher Guide And Answers Free and Paid eBooks
 - Redox Reactions Teacher Guide And Answers Public Domain eBooks
 - Redox Reactions Teacher Guide And Answers eBook Subscription Services
 - Redox Reactions Teacher Guide And Answers Budget-Friendly Options

6. Navigating Redox Reactions Teacher Guide And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Redox Reactions Teacher Guide And Answers Compatibility with Devices
 - Redox Reactions Teacher Guide And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox Reactions Teacher Guide And Answers
 - Highlighting and Note-Taking Redox Reactions Teacher Guide And Answers
 - Interactive Elements Redox Reactions Teacher Guide And Answers
8. Staying Engaged with Redox Reactions Teacher Guide And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox Reactions Teacher Guide And Answers
9. Balancing eBooks and Physical Books Redox Reactions Teacher Guide And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox Reactions Teacher Guide And Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Redox Reactions Teacher Guide And Answers
 - Setting Reading Goals Redox Reactions Teacher Guide And Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Redox Reactions Teacher Guide And Answers
 - Fact-Checking eBook Content of Redox Reactions Teacher Guide And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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