

Redox: IB Review Questions

1. Consider the standard electrode potentials of the following reactions:



What is the value of the cell potential (in volts) for the spontaneous reaction?

- A. +1.69 B. +1.39 C. +0.92 D. +0.62

2. Which factors affect the amount of metal formed during electrolysis?

- I. Charge on the metal ion
- II. Current
- III. Time

- A. I and II only B. I and III only
C. II and III only D. I, II and III

3. Consider the following reactions.



Which statement is correct?

- A. $\text{Cu}^{2+}(\text{aq})$ will oxidize both Mg(s) and Zn(s) .
B. Zn(s) will reduce both $\text{Cu}^{2+}(\text{aq})$ and $\text{Mg}^{2+}(\text{aq})$.
C. $\text{Mg}^{2+}(\text{aq})$ will oxidize both Cu(s) and Zn(s) .
D. Cu(s) will reduce both $\text{Mg}^{2+}(\text{aq})$ and $\text{Zn}^{2+}(\text{aq})$.

4. Consider the standard electrode potentials of the following reactions.



What is the value of the cell potential (in V) for the following reaction?



- A. -0.35 B. -1.15 C. +0.30 D. +0.35

Redox Ib Review Questions And Answers

National Cancer Institute (U.S.)



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