

Chem 12 Electrochemistry : Quiz 2 - 120

Part 1

1) A particular reaction has a predicted E° of -2.25 V. This indicates that the reaction would :

- a) occur rapidly b) produce energy c) go to completion
d) be non-spontaneous

2) Corrosion of a metal means that the metal :

- a) acts as a cathode b) undergoes oxidation
c) undergoes reduction d) acts as an oxidizing agent

3) Zinc metal is prepared by electrolysis as follows:



What is the minimum voltage necessary for the electrolysis reaction to proceed ?

- a) 2.12 b) 1.58 c) 1.96 d) 3.48

4) Which metal would provide the best cathodic protection for iron pipes ?

- a) Ag b) Sn c) Mg d) Cu

5) Which half-reaction best describes what is taking place at the anode during the electrolysis of a 1.0 M NaCl solution ?

- a) $\text{Na}^{+} + \text{e}^{-} \rightarrow \text{Na}$ b) $2\text{Cl}^{-} \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^{-}$
c) $2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-} \rightarrow \text{H}_2(\text{g})$ d) $\text{H}_2\text{O} \rightarrow 1/2 \text{O}_2(\text{g}) + 2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-}$

6) Which one of the following converts chemical energy to electrical energy ?

- a) Industrial production of aluminum b) Nickel plating
c) Discharging a "dry" cell d) Production of Na(s)

7) Which is the unit for the energy per coulomb of electrons ?

Redox And Electrochemistry Quiz Answers

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Competition for Electrons Thomas T. Mitsunaga, 1978 A First Course in Electrode Processes Derek Pletcher, 2009 This introduction highlights the importance of electrochemistry and its applications to the modern world and the future It emphasises understanding and avoids using many pages of complex equations It also describes the diverse applications of electrochemistry rather than focusing on analytical chemistry alone *Redox - Principles and Advanced Applications* , 19??

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