

Chapter 1

Problem Solutions

1.1

- (a) fcc: 8 corner atoms $\times 1/8 = 1$ atom
6 face atoms $\times 1/2 = 3$ atoms
Total of 4 atoms per unit cell

- (b) bcc: 8 corner atoms $\times 1/8 = 1$ atom
1 enclosed atom $= 1$ atom
Total of 2 atoms per unit cell

- (c) Diamond: 8 corner atoms $\times 1/8 = 1$ atom
6 face atoms $\times 1/2 = 3$ atoms
4 enclosed atoms $= 4$ atoms
Total of 8 atoms per unit cell

1.2

- (a) 4 Ga atoms per unit cell

$$\text{Density} = \frac{4}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ga} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

4 As atoms per unit cell, so that

$$\underline{\text{Density of As} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

- (b)

8 Ge atoms per unit cell

$$\text{Density} = \frac{8}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ge} = 4.44 \times 10^{22} \text{ cm}^{-3}}$$

1.3

- (a) Simple cubic lattice; $a = 2r$

$$\text{Unit cell vol} = a^3 = (2r)^3 = 8r^3$$

$$1 \text{ atom per cell, so atom vol.} = (1) \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{\left(\frac{4\pi r^3}{3} \right)}{8r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 52.4\%}$$

- (b) Face-centered cubic lattice

$$d = 4r = a\sqrt{2} \Rightarrow a = \frac{d}{\sqrt{2}} = 2\sqrt{2}r$$

$$\text{Unit cell vol} = a^3 = (2\sqrt{2}r)^3 = 16\sqrt{2}r^3$$

$$4 \text{ atoms per cell, so atom vol.} = 4 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{4 \left(\frac{4\pi r^3}{3} \right)}{16\sqrt{2}r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 74\%}$$

- (c) Body-centered cubic lattice

$$d = 4r = a\sqrt{3} \Rightarrow a = \frac{4}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{4}{\sqrt{3}}r \right)^3$$

$$2 \text{ atoms per cell, so atom vol.} = 2 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{2 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{4}{\sqrt{3}}r \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 68\%}$$

- (d) Diamond lattice

$$\text{Body diagonal} = d = 8r = a\sqrt{3} \Rightarrow a = \frac{8}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{8r}{\sqrt{3}} \right)^3$$

$$8 \text{ atoms per cell, so atom vol.} = 8 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{8 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{8r}{\sqrt{3}} \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 34\%}$$

1.4

From Problem 1.3, percent volume of fcc atoms is 74%. Therefore after coffee is ground,

$$\underline{\text{Volume} = 0.74 \text{ cm}^3}$$

Problem Physics Manual

Léo Pomar

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Table of Contents Problem Physics Manual

1. Understanding the eBook Problem Physics Manual
 - The Rise of Digital Reading Problem Physics Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Problem Physics Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Problem Physics Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Problem Physics Manual
 - Personalized Recommendations
 - Problem Physics Manual User Reviews and Ratings
 - Problem Physics Manual and Bestseller Lists
5. Accessing Problem Physics Manual Free and Paid eBooks
 - Problem Physics Manual Public Domain eBooks
 - Problem Physics Manual eBook Subscription Services
 - Problem Physics Manual Budget-Friendly Options
6. Navigating Problem Physics Manual eBook Formats

- ePub, PDF, MOBI, and More
- Problem Physics Manual Compatibility with Devices
- Problem Physics Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Problem Physics Manual
 - Highlighting and Note-Taking Problem Physics Manual
 - Interactive Elements Problem Physics Manual
- 8. Staying Engaged with Problem Physics Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Problem Physics Manual
- 9. Balancing eBooks and Physical Books Problem Physics Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Problem Physics Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Problem Physics Manual
 - Setting Reading Goals Problem Physics Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Problem Physics Manual
 - Fact-Checking eBook Content of Problem Physics Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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