

Write Titrating Volume (Spreads: 1.00)

Name _____

Acetic Acid (pKa = 4.75) titrated with NaOH

a. 0.100 M / volume 10.00 mL / pH 3.75 / 0.100 M

b. 0.100 M / volume 10.00 mL / pH 4.75 / 0.100 M

c. 0.100 M / pH 4.75 / compound that releases acetate
0.100 M / 10.00 mL / 0.100 M / 0.100 M

Acetic Acid: 1st, 2nd, 3rd, 4th, 5th, 6th, 7th, 8th, 9th, 10th

1.00 / 1.00 / 1.00 / 1.00 / 1.00 / 1.00 / 1.00 / 1.00 / 1.00 / 1.00

Species	Acid (pH, Base pH, or Salt pH)	Acids with M_1 ($M_1 = M_2$)	Conductivity (C or hC)
CH_3COOH	<u>4.75</u>	<u>1</u>	<u>1</u>
CH_3COO^-	<u>4.75</u>	<u>1</u>	<u>1</u>
$NaCH_3COO$	<u>4.75</u>	<u>1</u>	<u>1</u>
$NaCH_3COO^-$	<u>4.75</u>	<u>1</u>	<u>1</u>

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