Kedox Exercise (Part 2)

Chapter 20 Worksheet: Redox ANSWERS

 Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

2Sr + O₂ → 2SrO Sr⁰ to Sr²⁺; oxidized/reducing agent O⁰ to O²; reduced/ox. ag.

2. 2Li + S → Li₂S Li⁰ to Li^{1*}; oxidized/red. ag. S⁰ to S^{3*}; reduced/ox, ag.

2Cs + Br₂ - 2CsBr Cs⁶ to Cs¹⁺: oxidized/red. sg. Br⁰ to Br¹⁺; reduced/ox. ag.

3Mg + N₂ → Mg₃N₂ Mg⁰ to Mg²⁺; oxidized/red. ag. N⁰ to N³⁺; reduced/ox. ag.

4Fe + 3O₂ → 2Fe₂O₃ Fe⁶ to Fe^{3*}; oxidized/red. ag. O⁶ to O^{1*}; reduced/ox. ag.

Cl₂ + 2NaBr → 2NaCl + Br₂ Cl⁰ to Cl¹; reduced/ox. ag. Br¹ to Br⁰; oxidized/red. ag.

Si + 2F₂ → SiF₄ Si⁰ to Si⁴⁰; exiding freed, ag F⁰ to F¹; reduced/ox, ag.

Mg + 2HCl → MgCl₂ + H₂ Mg⁰ to Mg²⁺; oxidized/red. ag. H¹⁺ to H⁰; reduced/o.a.

10. 2Na + 2H₂O → 2NaOH + H₂ Na⁰ to Na^{1*}; oxidized/r.a. H^{1*} to H⁰; reduced/o.a.

Give the oxidation number of each kind of atom or ion.

a. sulfate b. Sn c. S²⁻ d. Fe³⁺ e. Sn⁴⁺ f. nitrate g. ammonium 2- 0 2- 3+ 4+ 1- 1+

12. Calculate the oxidation number of chromium in each of the following.

a. Cr_2O_3 b. $Na_2Cr_2O_7$ c. $CrSO_4$ d. chromate e. dichromate 3+ 6+ 2+ -7+ 6+

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

a. C + H₂SO₄ → CO₂ + SO₂ + H₂O C⁰ to C⁴⁺; oxidized S⁶⁺ to S⁴⁺; reduced

b. HNO₃ + HI -- NO + I₂ + H₂O N⁵⁺ to N²⁺; reduced I¹⁻ to I⁶; exidized

c. KMnO₄ + HCl -- MnCl₂ + Cl₂ + H₂O + KCl Mn² to Mn²; reduced Cl to Cl condized

d. Sb + HNO₃ \longrightarrow Sb₂O₃ + NO + H₂O Sb⁰ to Sb³⁺; oxidized N⁵⁺ to N²⁺; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

a. oxidizing agent: sulfur reducing agent: carbon

b. oxidizing agent: nitrogen reducing agent: iodine

oridizing agent, manganese reducing agent: chloride.

d. oxidizing agent: nitrogen reducing agent: antimony

Redox Problems And Answers

Liying Dong



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