

## SALT ANALYSIS



### SCHEME OF ANALYSIS

| EXPERIMENT                                                                                                                                      | OBSERVATION                                                                                                              | INFERENCE                                                                                                                                               |
|-------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>PRELIMINARY EXAMINATIONS :</b>                                                                                                               |                                                                                                                          |                                                                                                                                                         |
| 1) Noted the colour                                                                                                                             | Colourless                                                                                                               | Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$                                             |
|                                                                                                                                                 | Blue                                                                                                                     | Pre. of $\text{Cu}^{2+}$                                                                                                                                |
|                                                                                                                                                 | Brown                                                                                                                    | Pre. of $\text{Fe}^{3+}$                                                                                                                                |
|                                                                                                                                                 | Green                                                                                                                    | Pre. of $\text{Ni}^{2+}$                                                                                                                                |
|                                                                                                                                                 | Pale pink                                                                                                                | Pre. of $\text{Mn}^{2+}$                                                                                                                                |
|                                                                                                                                                 | Pink                                                                                                                     | Pre. of $\text{Co}^{2+}$                                                                                                                                |
| 2) Noted the solubility<br>a) In water<br>b) In dil.HCl                                                                                         | a) Soluble and no ppt with $\text{Na}_2\text{CO}_3$<br>b) Insoluble                                                      | a) Pre. of $\text{NH}_4^+$<br>b) Pre. of $\text{Pb}^{2+}$                                                                                               |
| 3) Flame Test :<br>A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.            | a) Blue surrounded by green<br>b) Brick red<br>c) Crimson red<br>d) Bluish white<br>e) Grassy green on prolonged heating | a) Pre. of $\text{Cu}^{2+}$<br>b) Pre. of $\text{Ca}^{2+}$<br>c) Pre. of $\text{Sr}^{2+}$<br>d) Pre. of $\text{Pb}^{2+}$<br>e) Pre. of $\text{Ba}^{2+}$ |
| 4) Ash Test :<br>A little of the salt is heated with con. $\text{HNO}_3$ and cobalt nitrate solution. A filter paper is dipped in it and burnt. | a) Green ash<br>b) Blue tinted ash<br>c) Pale pink ash                                                                   | a) Pre. of $\text{Zn}^{2+}$<br>b) Pre. of $\text{Al}^{3+}$<br>c) Pre. of $\text{Mg}^{2+}$                                                               |
| <b>SYSTEMATIC ANALYSIS OF ANIONS :</b>                                                                                                          |                                                                                                                          |                                                                                                                                                         |
| 1) Dil. $\text{H}_2\text{SO}_4$ test :<br>To the salt add dil. $\text{H}_2\text{SO}_4$ , warm if required.                                      | a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky.<br>b) Smell of vinegar        | a) Pre. of carbonates<br>b) Pre. of acetate                                                                                                             |

# Salt Analysis Manual For P G Course

**LM Reder**



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