

8 • Electron Configuration & the Periodic Table

AP QUESTION

- (a) Write the ground state electron configuration (short form or long form) for an arsenic atom.

As $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$

or As $[Ar] 4s^2 3d^{10} 4p^3$ or $[Ar] 3d^{10} 4s^2 4p^3$

- (b) Give one permissible set of four quantum numbers for each of the outermost electrons in a single As atom when it is in its ground state.

$4p$

4	1	-1	$+\frac{1}{2}$
---	---	----	----------------

$4p$

4	1	0	$+\frac{1}{2}$
---	---	---	----------------

$4p$

4	1	+1	$+\frac{1}{2}$
---	---	----	----------------

$(\frac{0}{-\frac{1}{2}})$

$4s$

4	0	0	$+\frac{1}{2}$
---	---	---	----------------

$4s$

4	0	0	$-\frac{1}{2}$
---	---	---	----------------

- (c) Is an isolated arsenic atom in the ground state paramagnetic or diamagnetic? Explain briefly.

Paramagnetic

$4p^3 = \uparrow \uparrow \uparrow$

half-filled orbitals = paramagnetic

- (d) What ion would As be expected to form? As³⁻
Write the electron configuration for the arsenic ion.

As $[Ar] 4s^2 3d^{10} 4p^6$ or As $[Ar] 3d^{10} 4s^2 4p^6$

As gains electrons and has the same electron configuration as Kr.

3-	2-	1-	He
N	O	F	Ne
P	S	Cl	Ar
As	Se	Br	Kr

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