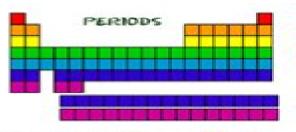
Unit 3 Notes: Periodic Table Notes

- John Newlands proposed an organization system based on increasing atomic mass in 1864.
- He noticed that both the chemical and physical properties repeated every 8
 elements and called this the Law of Octaves
- In 1869 both Lothar Meyer and Dmitri Mendeleev showed a connection between atomic mass and an element's properties.
- Mendeleev published first, and is given credit for this.
- He also noticed a periodic pattern when elements were ordered by increasing ____Atomic Mass
- By arranging elements in order of increasing atomic mass into columns, Mendeleev created the first Periodic Table.
- This table also predicted the existence and properties of undiscovered elements.
- After many new elements were discovered, it appeared that a number of elements were out of order based on their <u>Properties</u>
- By rearranging the elements based on ______ Atomic Number ____, the problems with the Periodic Table were corrected.
- This new arrangement creates a periodic repetition of both physical and chemical properties known as the _____Periodic Law____.



Periods are the ____Rows____

GROUPS

Groups/Families are the Columns

Valence electrons across a period are in the same energy level There are equal numbers of valence electrons in a group.

When elements are arranged in order of increasing <u>Atomic Number</u>,
 there is a periodic repetition of their <u>physical</u> and <u>chemical properties</u>

Periodic Trends Packet Key

Minjie Lin

Periodic Trends Packet Key:

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 - Chapter 4: Periodic Trends Packet Key in Specific Contexts
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Table of Contents Periodic Trends Packet Key

1. Understanding the eBook Periodic Trends Packet Key

- The Rise of Digital Reading Periodic Trends Packet Key
- Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Packet Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Packet Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Packet Key
 - Personalized Recommendations
 - Periodic Trends Packet Key User Reviews and Ratings
 - Periodic Trends Packet Key and Bestseller Lists
- 5. Accessing Periodic Trends Packet Key Free and Paid eBooks
 - Periodic Trends Packet Key Public Domain eBooks
 - Periodic Trends Packet Key eBook Subscription Services
 - Periodic Trends Packet Key Budget-Friendly Options
- 6. Navigating Periodic Trends Packet Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Packet Key Compatibility with Devices
 - Periodic Trends Packet Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Packet Key
 - Highlighting and Note-Taking Periodic Trends Packet Key
 - Interactive Elements Periodic Trends Packet Key
- 8. Staying Engaged with Periodic Trends Packet Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Packet Key

- 9. Balancing eBooks and Physical Books Periodic Trends Packet Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Packet Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Packet Key
 - Setting Reading Goals Periodic Trends Packet Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Packet Key
 - Fact-Checking eBook Content of Periodic Trends Packet Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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