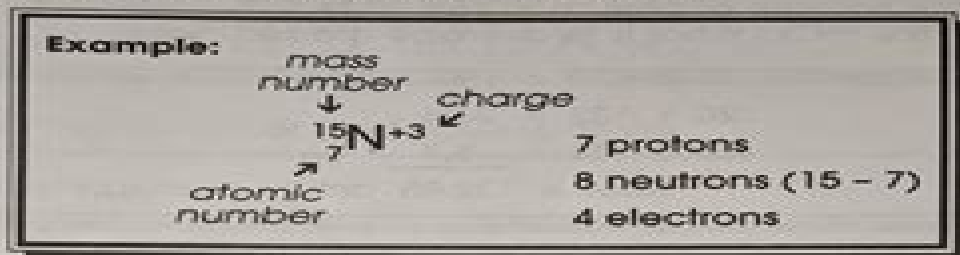


ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.



Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H ⁺						
$^{12}_6\text{C}$						
$^7_3\text{Li}^+$						
$^{35}_{17}\text{Cl}^-$						
$^{39}_{19}\text{K}$						
$^{24}_{12}\text{Mg}^{2+}$						
As ³⁻						
Ag						
Ag ⁺¹						
S ⁻²						
U						

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Mark Steyvers



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Table of Contents Periodic Table Worksheet 5 Answer Key Chemistry If8766

1. Understanding the eBook Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - The Rise of Digital Reading Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Personalized Recommendations
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 User Reviews and Ratings
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 and Bestseller Lists
5. Accessing Periodic Table Worksheet 5 Answer Key Chemistry If8766 Free and Paid eBooks
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 Public Domain eBooks
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 eBook Subscription Services
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 Budget-Friendly Options
6. Navigating Periodic Table Worksheet 5 Answer Key Chemistry If8766 eBook Formats

- ePub, PDF, MOBI, and More
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 Compatibility with Devices
 - Periodic Table Worksheet 5 Answer Key Chemistry If8766 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Highlighting and Note-Taking Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Interactive Elements Periodic Table Worksheet 5 Answer Key Chemistry If8766
 8. Staying Engaged with Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Worksheet 5 Answer Key Chemistry If8766
 9. Balancing eBooks and Physical Books Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Worksheet 5 Answer Key Chemistry If8766
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Setting Reading Goals Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Fact-Checking eBook Content of Periodic Table Worksheet 5 Answer Key Chemistry If8766
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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