

$$r^3 \propto A$$

$$r \propto A^{1/3}$$

- (b) Using the equation given in the problem, we can first solve for the radius of the lithium nucleus and then solve for its volume.

$$r = r_0 A^{1/3}$$

$$r = (1.2 \times 10^{-15} \text{ m})(7)^{1/3}$$

$$r = 2.3 \times 10^{-15} \text{ m}$$

$$V = \frac{4}{3} \pi r^3$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi (2.3 \times 10^{-15} \text{ m})^3 = 5.1 \times 10^{-44} \text{ m}^3$$

- (c) In part (b), the volume of the nucleus was calculated. Using the radius of a Li atom, the volume of a Li atom can be calculated.

$$V_{\text{atom}} = \frac{4}{3} \pi r^3 = \frac{4}{3} \pi (152 \times 10^{-12} \text{ m})^3 = 1.47 \times 10^{-29} \text{ m}^3$$

The fraction of the atom's volume occupied by the nucleus is:

$$\frac{V_{\text{nucleus}}}{V_{\text{atom}}} = \frac{5.1 \times 10^{-44} \text{ m}^3}{1.47 \times 10^{-29} \text{ m}^3} = 3.5 \times 10^{-15}$$

Yes, this calculation shows that the volume of the nucleus is *much, much* smaller than the volume of the atom, which supports Rutherford's model of an atom.

- 2.99** Two different structural formulas for the molecular formula $\text{C}_2\text{H}_6\text{O}$ are:



In the second hypothesis of Dalton's Atomic Theory, he states that in any compound, the ratio of the number of atoms of any two of the elements present is either an integer or simple fraction. In the above two compounds, the ratio of atoms is the same. This does not necessarily contradict Dalton's hypothesis, but Dalton was not aware of chemical bond formation and structural formulas.

- 2.100** (a)

Ethane
2.65 g C
0.665 g H

Acetylene
4.56 g C
0.383 g H

Let's compare the ratio of the hydrogen masses in the two compounds. To do this, we need to start with the same mass of carbon. If we were to start with 4.56 g of C in ethane, how much hydrogen would combine with 4.56 g of carbon?

$$0.665 \text{ g H} \times \frac{4.56 \text{ g C}}{2.65 \text{ g C}} = 1.14 \text{ g H}$$

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