

If 16.0 g of CaCO_3 are decomposed and 7.54 g of CaO are obtained, what is the percent yield of CaO obtained? \leftarrow A.Y



$$\text{Percent yield} = \frac{\text{Actual yield}}{\text{theoretical yield}} \times 100$$

$$16\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.1\text{g CaCO}_3} \times \frac{1\text{mol CaO}}{1\text{mol}}$$

Percent Actual And Theoretical Yield Key

Graham P. Bunn



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Chemistry: A Basic Introduction G. Tyler Miller, Jr., 1981 [Chemistry Textbook for College and University USA](#) Ibrahim Sikder, 2023-06-04

Chemistry Textbook USA *AP Chemistry For Dummies* Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13

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