

## Percent Yield Practice

1. Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.



$$\frac{24.8 \text{ g CaCO}_3}{1} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaCO}_3} \cdot \frac{56.08 \text{ g CaO}}{1 \text{ mol CaO}} = 13.9 \text{ g CaO} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{13.1}{13.9} \cdot 100 = \boxed{94.24\%} \Rightarrow \boxed{94.2\%}$$

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?



$$\frac{1.87 \text{ g Al}}{1} \cdot \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \cdot \frac{159.62 \text{ g CuSO}_4}{1 \text{ mol CuSO}_4} = 16.6 \text{ g CuSO}_4 \quad \text{Used}$$

$$\frac{9.65 \text{ g CuSO}_4}{1} \cdot \frac{1 \text{ mol CuSO}_4}{159.62 \text{ g CuSO}_4} \cdot \frac{3 \text{ mol Cu}}{3 \text{ mol CuSO}_4} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 3.84 \text{ g Cu}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{3.65}{3.84} \cdot 100 = \boxed{95.05\%} \quad \text{TY} \rightarrow \boxed{95.1\%}$$

3. If you start with 5.0 g of  $\text{C}_3\text{H}_8$  and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?



$$\frac{5.0 \text{ g C}_3\text{H}_8}{1} \cdot \frac{1 \text{ mol C}_3\text{H}_8}{44.11 \text{ g C}_3\text{H}_8} \cdot \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 8.2 \text{ g H}_2\text{O} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100$$

$$\frac{\text{PY} \cdot \text{TY}}{100} = \text{AY} = \frac{75 \cdot 8.2}{100} = \boxed{6.2 \text{ g H}_2\text{O}} \quad \text{AY}$$

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?



$$\frac{20 \text{ g CaO}}{1} \cdot \frac{1 \text{ mol CaO}}{56.08 \text{ g CaO}} \cdot \frac{1 \text{ mol Na}_2\text{O}}{1 \text{ mol CaO}} \cdot \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{20 \text{ g Na}_2\text{O}}$$

# Percent Yield Practice Problems Answer Key

**Lei Huang**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Percent Yield Practice Problems Answer Key**

1. Understanding the eBook Percent Yield Practice Problems Answer Key
  - The Rise of Digital Reading Percent Yield Practice Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Percent Yield Practice Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Percent Yield Practice Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Yield Practice Problems Answer Key
  - Personalized Recommendations
  - Percent Yield Practice Problems Answer Key User Reviews and Ratings
  - Percent Yield Practice Problems Answer Key and Bestseller Lists
5. Accessing Percent Yield Practice Problems Answer Key Free and Paid eBooks

- Percent Yield Practice Problems Answer Key Public Domain eBooks
  - Percent Yield Practice Problems Answer Key eBook Subscription Services
  - Percent Yield Practice Problems Answer Key Budget-Friendly Options
6. Navigating Percent Yield Practice Problems Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Percent Yield Practice Problems Answer Key Compatibility with Devices
    - Percent Yield Practice Problems Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Percent Yield Practice Problems Answer Key
    - Highlighting and Note-Taking Percent Yield Practice Problems Answer Key
    - Interactive Elements Percent Yield Practice Problems Answer Key
  8. Staying Engaged with Percent Yield Practice Problems Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Percent Yield Practice Problems Answer Key
  9. Balancing eBooks and Physical Books Percent Yield Practice Problems Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Percent Yield Practice Problems Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Percent Yield Practice Problems Answer Key
    - Setting Reading Goals Percent Yield Practice Problems Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Percent Yield Practice Problems Answer Key
    - Fact-Checking eBook Content of Percent Yield Practice Problems Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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