

- Q6 When water was added to the anhydrate of  $\text{CuSO}_4(x)$ , did the reaction proceed in the direction of the reactants or products?

## B.2 Reaction of Copper(II) Ion and Hydroxide

|                                               | Test tube 1       | Test tube 2      |
|-----------------------------------------------|-------------------|------------------|
| Initial color of $\text{CuCl}_2$ solution (1) | light blue        | light blue       |
| Initial appearance $\text{Cu(OH)}_2(x)$ (2)   | clear             | clear            |
| Change in appearance (3)                      | cloudy            | <del>clear</del> |
| Change in appearance (4)                      | <del>cloudy</del> | clear            |

- Q7 In test tube 1, did the addition of  $\text{OH}^-$  cause the reaction to go in the direction of the reactants or products? Explain.
- Q8 In test tube 2, did the addition of  $\text{H}_3\text{O}^+$  cause the reaction to go in the direction of the reactants or products? Explain.

## C. Changing Equilibrium Conditions: Le Châtelier's Principle

### C.1 Effect of Changing Reactant Concentration

| Test tube                   | Color      | Is color deeper or lighter red compared to the reference? | $[\text{FeSCN}^{2+}]$ increases or decreases? |
|-----------------------------|------------|-----------------------------------------------------------|-----------------------------------------------|
| 1. Reference                | red-orange | <del>deeper/lighter</del>                                 | <del>increases/decreases</del>                |
| 2. $\text{Fe}^{3+}$ added   | dark red   | deeper                                                    | increased                                     |
| 3. $\text{SCN}^-$ added     | blood red  | deeper                                                    | increased                                     |
| 4. $\text{Fe}^{3+}$ removed | orange     | lighter                                                   | decreased                                     |

# Reaction Rates And Equilibrium Lab Answers

**RJ Alexander**



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