

# Graphing Periodic Trends Activity

Name: \_\_\_\_\_ Period: \_\_\_\_\_

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	<p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p>						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	<p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>						
Rb							
Cs							

# Periodic Trends Extension Questions Answer Key

**Thomas Griffiths**



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## **Table of Contents Periodic Trends Extension Questions Answer Key**

1. Understanding the eBook Periodic Trends Extension Questions Answer Key
  - The Rise of Digital Reading Periodic Trends Extension Questions Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Extension Questions Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Extension Questions Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Extension Questions Answer Key
  - Personalized Recommendations
  - Periodic Trends Extension Questions Answer Key User Reviews and Ratings
  - Periodic Trends Extension Questions Answer Key and Bestseller Lists
5. Accessing Periodic Trends Extension Questions Answer Key Free and Paid eBooks
  - Periodic Trends Extension Questions Answer Key Public Domain eBooks
  - Periodic Trends Extension Questions Answer Key eBook Subscription Services
  - Periodic Trends Extension Questions Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Extension Questions Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Periodic Trends Extension Questions Answer Key Compatibility with Devices
- Periodic Trends Extension Questions Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Extension Questions Answer Key
  - Highlighting and Note-Taking Periodic Trends Extension Questions Answer Key
  - Interactive Elements Periodic Trends Extension Questions Answer Key
- 8. Staying Engaged with Periodic Trends Extension Questions Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Extension Questions Answer Key
- 9. Balancing eBooks and Physical Books Periodic Trends Extension Questions Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Extension Questions Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Extension Questions Answer Key
  - Setting Reading Goals Periodic Trends Extension Questions Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Extension Questions Answer Key
  - Fact-Checking eBook Content of Periodic Trends Extension Questions Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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