

Periodic Trends: Straw Lab

Objective: To create a visual representation of the following periodic trends: atomic size, ionization energy and electronegativity.

Atomic Radius: the size of an atom measured in either nanometers (nm) or angstroms (Å).

Ionization Energy: the energy needed to remove an electron from an atom measured in kilojoules per mole (kJ/mol) of an atom.

Electronegativity: the relative ability of an atom to attract electrons towards itself.

Materials: For this lab you will need a sheet with 3 small periodic tables, straws, scissors, a metric ruler, your notebook and access to a hot glue gun.

Procedure:

- 1) Draw and complete the data table for each trend in your lab notebook. Each person should have their own data tables in their own notebooks!
- 2) For each trend, calculate the scale factor according to the instructions. Using the scale factor, measure the straw and cut it to the appropriate size.
- 3) You will then use the glue gun to glue the straw to the periodic table in the box that represents the element's location on the periodic table.
- 4) Answer the questions for each trend in your lab notebook. Everyone should have their own answers in their own notebook!

Atomic Radius:

The radius of the following atoms has been estimated to have the following values:

Name	Symbol	Atomic Number	Size of Radius (nm)	Calculated Straw Length (cm)
Fluorine			0.064	
Chlorine			0.0999	
Bromine			0.114	
Iodine			0.1133	
Sodium			0.186	
Magnesium			0.160	
Aluminum			0.143	
Silicon			0.117	

Periodic Trends Straw Lab Answer Key

Richard Bailey



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