Periodic Trends Worksheet

Name: _____ ID: ____

Directions: Use your notes to answer the following questions.

Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Ans: 0, C, Al, K

Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Ans: Ne, Al, S, O

(3) Why does fluorine have a higher ionization energy than iodine?

Ans: Fluorine has a smaller atomic radius, so the proton can exert a greater pull on electron (Valence shell is close)

Why do elements in the same family generally have similar properties?

Ans: They have the same valence shell electron arrangement.

- 5 Indicate whether the following properties increase or decrease from left to right across the periodic table.
 - atomic radius (excluding noble gases) Decrease
 - b. first ionization energy Increase
 - electronegativity Increase
- (6) What trend in atomic radius occurs across the periodic table? What causes this trend?

Ans: Atomic radius increases down a group because energy levels (shells) are added.

What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ans: Ionization energy increases across a period because as elements become less metallic, it requires more energy to remove an electron.

8 Circle the atom in each pair that has the largest radius.

a. (Al) or B d. (O) or F

b. (Na) or Al e. (Br) or Cl

c. S or O f. Mg or Ca

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