Lab 6: CHEMICAL EQUILIBRIUM LAB

Introduction: In this experiment you will start with a reaction that is at equilibrium. You will then change the concentrations of various ions that are present and note the effect on the state of equilibrium. The equilibrium reaction chosen has a colored product, so changes in its concentration can be readily detected by a corresponding change in color intensity.

Procedure:

	Pour about	25 mL of 0.0020	M potassium thiocyanate solution,	KSCN,
	and 25 mL	of distilled water	r into a beaker.	

- a. What is the color of this solution? _____
- b. What are the 2 ions in this solution? _____
- Look at the bottle of 0.20M iron (III)nitrate solution, Fe(NO₃)₃ at your lab station.
 - a. What is the color of this solution? _____
 - b. What are the 2 ions in this solution? _____
- Add 5 drops of the Fe(NO₃)₃ solution to the beaker containing the
 potassium thiocyanate solution, KSCN. Gently swirl the mixture.
 What is the color of this solution? ______
- 4. There are now 4 ions in this solution. What are the two combinations of ions that could account for the color in step three?
- 5. At your lab station there is a solution of potassium nitrate KNO₃. What is the color of this solution? ______
 What are the two ions in this solution? ______
- Based on your observations in step 5 and your answer to question 4, what are the two ions that must be responsible for the color you observed in step 5?

Observing Chemical Equilibrium Lab 46 Answers

SJ Ball

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