

Answer KEY

Unit 3 Outline/ Study Guide - Electron Configuration and Periodic Trends

Essential Skills/ State Standards:

1. Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
2. Students know how to use the periodic table to identify metals, semimetals, nonmetals, halogens, alkali metals, alkaline earth metals, and transition metals.
3. Students know how to use the periodic table to identify the following trends:
 ionization energy, electronegativity, and the relative sizes of ions and atoms
 (this includes knowing the definitions of ionization energy, electronegativity, cation, and anion.)

Study Guide - (use your notes and worksheets for a more complete review)

1. Identify the following things on the periodic table:

-metals	-alkali metals
-metalloids (semi-metals)	-alkaline earth metals
-transition metals	-halogens
-nonmetals	-noble gases
-Periods	-Groups/ families
Identify groups with the following numbers of valence electrons: 1, 2, 3, 4, 5, 6, 7, 8	

2. Why do the elements in the groups (columns) behave so similar to each other even if they have great differences in their number of protons? **same number of valence electrons and similar chemical properties.**
3. What is the connection between the following terms: **octet rule, valence electrons, and ions?**
In order to achieve the octet rule, atoms gain or lose their valence electrons. They become charged ions when this occurs.
4. a. List the elements that are exceptions to the octet rule b) Why don't these atoms achieve the octet rule?
 a. Small atoms are happy with 2 electrons (ex: H, He, Li, Be, B)
 b. because the 1s energy level only holds 2 electrons

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