

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

1	2											13	14	15	16	17	18
1 H 1.01 Hydrogen	2 He 4.003 Helium											3 B 10.81 Boron	4 C 12.01 Carbon	5 N 14.01 Nitrogen	6 O 16.00 Oxygen	7 F 19.00 Fluorine	8 Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.012 Beryllium											9 Al 26.98 Aluminum	10 Si 28.09 Silicon	11 P 30.97 Phosphorus	12 S 32.06 Sulfur	13 Cl 35.45 Chlorine	14 Ar 39.95 Argon
5 Na 22.99 Sodium	6 Mg 24.31 Magnesium											15 Ga 69.72 Gallium	16 Ge 72.64 Germanium	17 As 74.92 Arsenic	18 Se 78.96 Selenium	19 Br 79.90 Bromine	20 Kr 83.80 Krypton
7 K 39.10 Potassium	8 Ca 40.08 Calcium											21 In 114.82 Indium	22 Sn 118.71 Tin	23 Sb 121.76 Antimony	24 Te 127.60 Tellurium	25 I 126.91 Iodine	26 Xe 131.29 Xenon
9 Rb 85.47 Rubidium	10 Sr 87.62 Strontium											27 Hg 200.59 Mercury	28 Tl 204.38 Thallium	29 Pb 207.2 Lead	30 Bi 208.98 Bismuth	31 Po 209 Polonium	32 At 210 Astatine
11 Cs 132.91 Cesium	12 Ba 137.33 Barium											33 Nb 92.91 Niobium	34 Mo 95.94 Molybdenum	35 Tc 98 Technetium	36 Ru 101.07 Ruthenium	37 Rh 102.91 Rhodium	38 Pd 106.42 Palladium
13 Fr 223 Francium	14 Ra 226 Radium											39 Y 88.91 Yttrium	40 Zr 91.22 Zirconium	41 Nb 92.91 Niobium	42 Mo 95.94 Molybdenum	43 Tc 98 Technetium	44 Ru 101.07 Ruthenium
												45 Rh 102.91 Rhodium	46 Pd 106.42 Palladium	47 Ag 107.87 Silver	48 Cd 112.41 Cadmium	49 In 114.82 Indium	50 Sn 118.71 Tin
												51 Cs 132.91 Cesium	52 Ba 137.33 Barium	53 La 138.91 Lanthanum	54 Ce 140.12 Cerium	55 Pr 140.91 Praseodymium	56 Nd 144.24 Neodymium
												57 Eu 151.96 Europium	58 Gd 157.25 Gadolinium	59 Tb 158.93 Terbium	60 Dy 162.50 Dysprosium	61 Ho 164.93 Holmium	62 Er 167.26 Erbium
												63 Lu 174.96 Lutetium	64 Hf 178.49 Hafnium	65 Ta 180.95 Tantalum	66 W 183.84 Tungsten	67 Re 186.21 Rhenium	68 Os 190.23 Osmium
												69 Ir 192.22 Iridium	70 Pt 195.08 Platinum	71 Au 196.97 Gold	72 Hg 200.59 Mercury	73 Tl 204.38 Thallium	74 Pb 207.2 Lead
												75 Bi 208.98 Bismuth	76 Po 209 Polonium	77 At 210 Astatine	78 Rn 222 Radon	79 Fr 223 Francium	80 Ra 226 Radium
												81 Ac 227 Actinium	82 Th 232.04 Thorium	83 Pa 231.04 Protactinium	84 U 238.03 Uranium	85 Np 237.05 Neptunium	86 Pu 244.06 Plutonium
												87 Am 243.06 Americium	88 Cm 247.07 Curium	89 Bk 247.07 Berkelium	90 Cf 251.08 Californium	91 Es 252.08 Einsteinium	92 Fm 257.10 Fermium
												93 Md 258.10 Mendelevium	94 No 259 Nobelium	95 Lr 260.10 Lawrencium	96 103 Dubnium	97 104 Rutherfordium	98 105 Dubnium
												99 106 Seaborgium	100 107 Bohrium	101 108 Hassium	102 109 Meitnerium	103 110 Darmstadtium	104 111 Roentgenium
												105 112 Copernicium	106 113 Nihonium	107 114 Flerovium	108 115 Moscovium	109 116 Livermorium	110 117 Tennessine
												111 118 Oganesson	112 119 Ununennium	113 120 Ubnium	114 121 Ubuium	115 122 Ubnium	116 123 Ubnium

CLASSIFICATION

Metals	Nonmetals
Active Pre-transition Alkali earth Alkaline	Corrosive
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals gas

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals, those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalyst properties. Very weakly electropositive; inert to acids.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and tellurates.

Metals

- Metals**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionization energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
B and Si are colorless gases, C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionization energies; low to high electron affinities; moderate to high electronegativity values; relatively poor to moderately strong oxidizing agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (P₂O, CO, NO, N₂O).
- Group 1**
The alkali metals, and the alkaline earth metals (Li, Be, Ca, Sr, Ba, Ra, and Fr) (except Fr, the most stable form in ambient conditions).
- Group 2**
The alkali metals, and the alkaline earth metals (Li, Be, Ca, Sr, Ba, Ra, and Fr) (except Fr, the most stable form in ambient conditions).
- Group 3**
The alkali metals, and the alkaline earth metals (Li, Be, Ca, Sr, Ba, Ra, and Fr) (except Fr, the most stable form in ambient conditions).
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- Group 18**
The alkali metals, and the alkaline earth metals (Li, Be, Ca, Sr, Ba, Ra, and Fr) (except Fr, the most stable form in ambient conditions).

Nonmetals

- Corrosive nonmetals**
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionization energies, electron affinities, and electronegativity values, and be relatively strong oxidizing agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesized in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answers

Raffaela Di Napoli

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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- Chapter 2: Essential Elements of Periodic Table Guide Answers
- Chapter 3: Periodic Table Guide Answers in Everyday Life
- Chapter 4: Periodic Table Guide Answers in Specific Contexts
- Chapter 5: Conclusion

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