

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1) 10.0 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.0 grams. The actual yield was 23.0 grams. Calculate the percent yield of sodium chloride.

- 2) 4.0 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 11.0 grams. Calculate the percent yield of carbon dioxide.

- 3) 1.0 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 1.5 grams. Calculate the percent yield of zinc sulfide.

- 4) 1.0 grams of phosphorus reacts with excess chlorine.



The theoretical yield of phosphorus trichloride is 3.0 grams. Calculate the percent yield of phosphorus trichloride.

- 5) Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.0 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 6) The actual yield was 10.0 grams. Calculate the percent yield.

- 7) Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). 10.0 grams of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 8) The actual yield was 40.0 grams. Calculate the percent yield.

- 9) Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 10.0 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 10.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 14.69 \text{ g CaCl}_2 \end{array}$$

$$10.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 14.69 \text{ g CaCl}_2$$

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Molar Mass CaCl_2

$$\text{Ca} = 40.08$$

$$\text{Cl} = 35.45$$

$$\text{Cl} = 35.45 \times 2 = 70.90$$

$$\text{Total} = 40.08 + 70.90 = 110.98$$

$$\text{Molar Mass} = 110.98 \text{ g}$$

Molar Mass CaCO_3

$$\text{Ca} = 40.08$$

$$\text{C} = 12.01$$

$$\text{O} = 16.00 \times 3 = 48.00$$

$$\text{Total} = 40.08 + 12.01 + 48.00 = 100.09$$

$$\text{Molar Mass} = 100.09 \text{ g}$$

- 10) The actual yield was 15.0 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{15.0 \text{ g}}{14.69 \text{ g}} \times 100 = 102.11\% \text{ CaCl}_2$$

- 11) Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 10.0 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 10.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 41.20 \text{ g KI} \end{array}$$

Molar Mass KI

$$\text{K} = 39.10$$

$$\text{I} = 126.90$$

$$\text{Total} = 39.10 + 126.90 = 166.00$$

$$\text{Molar Mass} = 166.00 \text{ g}$$

ANSWER KEY
INCLUDED

%
YIELD

Percent Yield Example Answer Key

Liyong Dong

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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